

General Chemistry II Jasperse
Acid-Base Chemistry: Extra Practice Problems

General Types/Groups of problems:

Conceptual Questions: Acids, Bases, and Conjugates, Miscellaneous	p1	K_b and pK_b , Base Strength, and using K_b or pK_b to Calculate $[OH^-]$, pOH , pH , and/or $[H^+]$	p7-10
Recognizing Strong versus Weak Acids; Recognizing Basic versus Nonbasic	p3	Recognizing Acid/Base Properties when Ionics are Dissolved in Water	p11
pH Calculations: Relationships between pH and pOH	p4	Answers	p12
K_a : Sense + Calculations: Using K_a or pK_a to Calculate $[H^+]$ and/or pH; using pH to calculate K_a or pK_a	p5		

Conceptual Questions: Acids, Bases, and Conjugates, Miscellaneous

- In the Brønsted-Lowry definition of acids and bases, an acid _____.
 - is a proton donor.
 - is a proton acceptor.
 - forms stable hydrogen bonds.
 - breaks stable hydrogen bonds.
 - corrodes metals.
- In the Brønsted-Lowry definition of acids and bases, a base _____.
 - is a proton donor.
 - is a proton acceptor.
 - forms stable hydrogen bonds.
 - breaks stable hydrogen bonds.
 - corrodes metals.
- In the following reaction in aqueous solution, the acid reactant is _____ and its conjugate base product is _____.

$$CH_3COOH + NH_3 \rightleftharpoons CH_3COO^- + NH_4^+$$
 - CH_3COOH ; CH_3COO^-
 - CH_3COOH ; NH_4^+
 - NH_3 ; CH_3COO^-
 - NH_3 ; NH_4^+
 - CH_3COOH ; H_3O^+
- In the following reaction in aqueous solution, the acid reactant is _____, and its conjugate base product is _____.

$$CH_3NH_2 + HSO_4^- \rightleftharpoons CH_3NH_3^+ + SO_4^{2-}$$
 - CH_3NH_2 ; $CH_3NH_3^+$
 - CH_3NH_2 ; SO_4^{2-}
 - HSO_4^- ; $CH_3NH_3^+$
 - HSO_4^- ; SO_4^{2-}
 - HSO_4^- ; H_3O^+
- Which of the following is the conjugate acid of the hydrogen phosphate ion, HPO_4^{2-} ?
 - $H_2PO_4^-$
 - $H_2PO_4^-$
 - HPO_4^{2-}
 - PO_4^{3-}
 - H_3O^+

General Chemistry Acid Base Practice Problems

Rachel S Tattersall



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Table of Contents General Chemistry Acid Base Practice Problems

1. Understanding the eBook General Chemistry Acid Base Practice Problems
 - The Rise of Digital Reading General Chemistry Acid Base Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Acid Base Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Acid Base Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Acid Base Practice Problems
 - Personalized Recommendations
 - General Chemistry Acid Base Practice Problems User Reviews and Ratings
 - General Chemistry Acid Base Practice Problems and Bestseller Lists

5. Accessing General Chemistry Acid Base Practice Problems Free and Paid eBooks
 - General Chemistry Acid Base Practice Problems Public Domain eBooks
 - General Chemistry Acid Base Practice Problems eBook Subscription Services
 - General Chemistry Acid Base Practice Problems Budget-Friendly Options
6. Navigating General Chemistry Acid Base Practice Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Acid Base Practice Problems Compatibility with Devices
 - General Chemistry Acid Base Practice Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Acid Base Practice Problems
 - Highlighting and Note-Taking General Chemistry Acid Base Practice Problems
 - Interactive Elements General Chemistry Acid Base Practice Problems
8. Staying Engaged with General Chemistry Acid Base Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Acid Base Practice Problems
9. Balancing eBooks and Physical Books General Chemistry Acid Base Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Acid Base Practice Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Acid Base Practice Problems
 - Setting Reading Goals General Chemistry Acid Base Practice Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Acid Base Practice Problems
 - Fact-Checking eBook Content of General Chemistry Acid Base Practice Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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