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Glenn Fleishman



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off zero point is 2.0 M acetone and 15 ml of I_2 . Assume this is a concentration of x and h in the first case we

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web iodination of acetone advance study assignment 1 in a reaction involving the iodination of acetone the following volumes were used to make up the reaction mixture 10 ml 4.0 M acetone 10 ml 1.0 M HCl 10 ml 0.0050 M I_2 20 ml H_2O a) how many moles of acetone h and I_2 are in the reaction mixture moles acetone moles h moles I_2 b)

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web video answer there can be a 6.0 M solution in sodium hydroxide ensure it has been reduced to 400 liters so we have to find the moles in the solution the first thing

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web the iodination of acetone is convenient because the color of the iodine reactant is readily observable and the change in concentration of the iodine can be easily monitored by our spectrophotometer and the reaction proceeds quickly enough for multiple trials but slowly enough to obtain measurements

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web I the iodination of acetone the rate of this reaction is found to depend on the concentration of hydrogen ion in the solution as well as presumably on the concentrations of the two reactants

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