

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the best (but not all) of the information for this chapter is that.

Focus on this content, but make sure to review other topics: reactions, bonding, molecules, etc.

If you study this document and nothing else, you should at least be able to PASS this test.

Good luck with the exam, exam, exam application of this content, exam.

OBJECTIVES

- Compare with periodic table (modern/chemical/elemental/atomic/physical)
- Compare with periodic table (modern/chemical/elemental/atomic/physical) (P. 12)
- 5.1: Describe the organization of the periodic table: structure, properties, groups, and periods (P. 12)
- 5.2: Describe the periodic table (modern/chemical/elemental/atomic/physical) and structure, properties, groups, and periods (P. 12)
- 5.3: Identify elements: metals, nonmetals, metalloids, and noble gases; identify the representative groups and transition metals (groups 1-10, 11, 12, 13, 14, 15, 16, 17, 18) (P. 12)
- 5.4: Compare periodic trends of atomic size, ionization energy, and electronegativity (P. 12)

5.1: HISTORY OF THE PERIODIC TABLE

Introduction

- periodic law = the law that states that the repeating chemical and physical properties of elements change periodically with the atomic number of the element
- periodic table = an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column
- classification = a method of the periodic table of elements, which classifies elements into groups, periods, and blocks (P. 12)
- classification = a method of the periodic table of elements, which classifies elements into groups, periods, and blocks (P. 12)

Classification

- The periodic law states that the physical and chemical properties of the elements, and periodic properties, of the elements, change periodically with the atomic number of the element
- The periodic table is an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column/group/family
- The columns in the periodic table are referred to as groups

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

Introduction

- atomic number = one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
- atomic number = one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
- transition elements = one of the elements, that are not the noble gases, which have a partially filled d-orbital
- transition elements = one of the elements of the periodic table, which are not the noble gases, which have a partially filled d-orbital
- transition elements = one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine), halogens, which are not the noble gases

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