

Chemistry 105

Chapter 3 Review Problems

- How many molecules are present in 180 grams of water? (The atomic weights of H and O are 1 g/mol and 16 g/mol, respectively.)
 - 6.0×10^{24}
 - 10
 - 5.6×10^{-23}
 - 1.8×10^7
 - 180×10^{23}
- The atomic mass of sulfur is given as 32.07 g/mol. What is the mass of an average atom of sulfur?
 - 5.33×10^{-23} g
 - 32.07 g
 - 32.07×10^{-23} kg
 - 193×10^{-23} g
 - 6.02×10^{23} g
- What is the mass of C present in 60 g of C_2H_6 ?
 - 12 g
 - 24 g
 - 48 g
 - 60 g
 - 6 g
- What is the molar mass of C_2H_6O ?
 - 24 g/mol
 - 32 g/mol
 - 7 g/mol
 - 35 g/mol
 - 46 g/mol
- What is the correct percentage by mass of nitrogen in HCN?
 - 33 %
 - 14%
 - 27%
 - 52%
 - 98%
- What is the chief danger that is anticipated if CO_2 levels get too high in the atmosphere?
 - people will develop skin cancer from exposure to UV light
 - the atmosphere will develop an unpleasant odor
 - the temperature of the planet may climb by several degrees
 - plant life will be exterminated
 - The CO_2 will remove ozone from the stratosphere

Holt Chemistry Standardized Test Prep Chapter 3 Answer Key

A Loxley



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