

CHAPTER 11 REVIEW

Ions in Aqueous Solutions and Colligative Properties

SECTION REVIEW

IMPORT QUESTIONS Answer the following questions in the space provided.

1. Identify the ions represented under each of the descriptions on the left.

- | | |
|---|---------------------------------------|
| _____ an ionic compound that is quite soluble in water | (a) NaCl |
| _____ an ionic compound that is not very soluble in water | (b) NaNO_3 |
| _____ a molecular compound that contains no oxygen | (c) AgCl |
| _____ a molecular compound that does not contain an acid | (d) $\text{C}_2\text{H}_5\text{COOH}$ |

2. Consider separately separately-saturated solutions of various liquid solutes in water. In each case, the following statements:

- | | |
|-------|--|
| _____ | a. The change in the boiling point decreases with the density of the solid solute, with the increasing solute density having no effect. |
| _____ | b. The change in the boiling point decreases with the density of the solid solute, with the increasing solute density having no effect. |
| _____ | c. The change in the boiling point becomes greater as the concentration of the solute in solution increases decreases. |

3. a. Explain why compound A is soluble in water while compound B is not. The formulas of compounds A and B are:

b. Identify any species ions in the system from dissociation process.

c. Write the net ionic equation for the formation of water in a solution.

4. Explain why applying road salt helps melt ice on icy sidewalks faster than shoveling snow.

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