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Modern Chemistry • CHAPTER 7

**HOMEWORK 7-1**

(pp. 263–266)

**VOCABULARY**Write *true* or *false* for each statement.

1. A chemical formula indicates the relative number of molecules of each kind in a chemical compound.

\_\_\_\_\_

2. Monoatomic ions are ions formed from a single atom. \_\_\_\_\_

3. Monoatomic anions of an element are identified by the element's name. \_\_\_\_\_

4. A binary compound is made of two or more different elements. \_\_\_\_\_

5. The Stock system is a system for naming chemical elements and their ions. \_\_\_\_\_

**SKILL BUILDER**

Explain what each chemical formula tells you about the composition of the compound. The first one is done for you.



The subscript 3 refers to the number of oxygen atoms. There are one atom of sodium and one atom of nitrogen for every three atoms of oxygen.

**STANDARDIZED TEST PREP**

Circle the letter of the best answer.

1. Which statement is true?
- All main-group elements are able to form ions.
  - The total numbers of positive charges and negative charges must be equal in a binary compound.
  - The nonmetals of groups 15, 16, and 17 lose electrons to form cations.
  - Elements in the *d* block form ions of only one charge.
2. Why do many main-group elements lose or gain electrons to form ions?
- The elements are unstable and therefore lose electrons.
  - Their tendency is to maximize potential energy.
  - The elements form covalent bonds, sharing their electrons with other atoms.
  - Their tendency is to form a complete outermost octet in a noble-gas configuration.

Modern Chemistry

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