

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

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1. Understanding the eBook Holt Mcdougal Chemistry Chapter Review Answers
 - The Rise of Digital Reading Holt Mcdougal Chemistry Chapter Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Mcdougal Chemistry Chapter Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Mcdougal Chemistry Chapter Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Mcdougal Chemistry Chapter Review Answers

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 - Holt Mcdougal Chemistry Chapter Review Answers User Reviews and Ratings
 - Holt Mcdougal Chemistry Chapter Review Answers and Bestseller Lists
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 - Interactive Elements Holt Mcdougal Chemistry Chapter Review Answers
 8. Staying Engaged with Holt Mcdougal Chemistry Chapter Review Answers
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Mcdougal Chemistry Chapter Review Answers
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Mcdougal Chemistry Chapter Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Holt Mcdougal Chemistry Chapter Review Answers
 - Setting Reading Goals Holt Mcdougal Chemistry Chapter Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Holt Mcdougal Chemistry Chapter Review Answers

- Fact-Checking eBook Content of Holt Mcdougal Chemistry Chapter Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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