Questions and Problems

Q1 According to your graph, what is the relationship between the pressure and the volume of a gas?

As the volume of the gas increases the pressure of the gas decreases. They have inverse proportional relationship.

Q2 On your graph, what is the volume of the gas at a pressure of 760. mmHg?

* After plotting the data to excel, the graph produced a linear equation y = -28.873x + 1539.3 where y is the pressure and x is the volume of gas.

We need to find x at 760 mmHg.

Hence,

760 mmHg = -28.873x + 1539.8 → x = 26.99 mL

Therefore, at 760 mmHg the volume of the gas is 26.99 mL

Q3 On your graph, what is the pressure of the gas when the volume is 30.0 mL?

$$x = 30.0 \text{ mL}, y = ?$$

$$y = -28.873(30.0) + 1539.3$$

$$y = 673.1 \text{ mmHg}$$

Therefore, when the volume of the gas is 30.0 mL the pressure is 673.1 mmHg

Q4 A sample of helium has a volume of 325 mL and a pressure of 655 mmHg. What will be the pressure, in mmHg, if the sample of helium is compressed to 125 mL (T, n constant)?

$$\mathbf{P}_1\mathbf{V}_1 = \mathbf{P}_2\mathbf{V}_2$$

$$P_2 = (P_1V_1) \div V_2 = (655 \text{ mmHg x } 325 \text{ mL}) \div 125 \text{ mL} = 1,703 \text{ mmHg}$$

Holt Chemistry Section Review Answers Gases

Liying Dong

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