

Problem Solving *continued*

- 12.** Urea, $(\text{NH}_2)_2\text{CO}$, is an important fertilizer that is manufactured by the following reaction.



What volume of NH_3 at STP will be needed to produce 8.50×10^4 kg of urea if there is an 89.5% yield in the process?

- 13.** An obsolete method of generating oxygen in the laboratory involves the decomposition of barium peroxide by the following equation.



What mass of BaO_2 reacted if 265 mL of O_2 is collected by water displacement at 0.975 atm and $10.^\circ\text{C}$?

- 14.** It is possible to generate chlorine gas by dripping concentrated HCl solution onto solid potassium permanganate according to the following equation.



If excess HCl is dripped onto 15.0 g of KMnO_4 , what volume of Cl_2 will be produced? The Cl_2 is measured at 15°C and 0.959 atm.

- 15.** Ammonia can be oxidized in the presence of a platinum catalyst according to the following equation.



The NO that is produced reacts almost immediately with additional oxygen according to the following equation.



If 35.0 kL of oxygen at STP react in the first reaction, what volume of NH_3 at STP reacts with it? What volume of NO_2 at STP will be formed in the second reaction, assuming there is excess oxygen that was not used up in the first reaction?

- 16.** Oxygen can be generated in the laboratory by heating potassium chlorate. The reaction is represented by the following equation.



What mass of KClO_3 must be used in order to generate 5.00 L of O_2 , measured at STP?

- 17.** One of the reactions in the Solvay process is used to make sodium hydrogen carbonate. It occurs when carbon dioxide and ammonia are passed through concentrated salt brine. The following equation represents the reaction.



- a.** What volume of NH_3 at 25°C and 1.00 atm pressure will be required if 38 000 L of CO_2 , measured under the same conditions, react to form NaHCO_3 ?
- b.** What mass of NaHCO_3 can be formed when the gases in (a) react with NaCl ?

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1. Understanding the eBook Holt Chemistry Practice Problems
 - The Rise of Digital Reading Holt Chemistry Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Practice Problems
 - Personalized Recommendations
 - Holt Chemistry Practice Problems User Reviews and Ratings
 - Holt Chemistry Practice Problems and Bestseller Lists
5. Accessing Holt Chemistry Practice Problems Free and Paid eBooks
 - Holt Chemistry Practice Problems Public Domain eBooks
 - Holt Chemistry Practice Problems eBook Subscription Services
 - Holt Chemistry Practice Problems Budget-Friendly Options
6. Navigating Holt Chemistry Practice Problems eBook Formats

- ePub, PDF, MOBI, and More
 - Holt Chemistry Practice Problems Compatibility with Devices
 - Holt Chemistry Practice Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Practice Problems
 - Highlighting and Note-Taking Holt Chemistry Practice Problems
 - Interactive Elements Holt Chemistry Practice Problems
 8. Staying Engaged with Holt Chemistry Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Practice Problems
 9. Balancing eBooks and Physical Books Holt Chemistry Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Practice Problems
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Holt Chemistry Practice Problems
 - Setting Reading Goals Holt Chemistry Practice Problems
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Holt Chemistry Practice Problems
 - Fact-Checking eBook Content of Holt Chemistry Practice Problems
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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