

Electrochemistry: Unit Review

1. What is the oxidation number of chlorine in HClO_4 ?
1. -1 2. +1 3. +7 4. +8
2. What is the oxidation number of carbon in C_3H_8 ?
1. -6 2. -3 3. +3 4. +6
3. A battery is an example of:
1. An electrochemical cell 2. A electrolytic cell 3. A pressure cell
4. Given the reaction: $\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$ Which equation represents the oxidation that takes place?
1. $\text{Mg}^{+2} + 2e^- \rightarrow \text{Mg}$ 2. $\text{Mg} \rightarrow \text{Mg}^{+2} + 2e^-$
3. $\text{Cu}^{+2} + 2e^- \rightarrow \text{Cu}$ 4. $\text{Cu} \rightarrow \text{Cu}^{+2} + 2e^-$
5. In the following reaction: $\text{Fe} + \text{Zn}^{+2} \rightarrow \text{Fe}^{+2} + \text{Zn}$
A. Fe is reduced and Zn^{+2} is the oxidizing agent
B. Fe is reduced and Zn^{+2} is the reducing agent
C. Fe is oxidized and Zn^{+2} is the oxidizing agent
D. Fe is oxidized and Zn^{+2} is the reducing agent
6. In which species does oxygen have an oxidation number of -1?
A. H_2O 2. Na_2O_2 3. Al_2O_3 4. O_2
7. Elements that readily gain electrons tend to have
A. High ionization energies and high electronegativities
B. High ionization energies and low electronegativities
C. Low ionization energies and high electronegativities
D. Low ionization energies and low electronegativities
8. In an electrochemical cell, the cathode is
1. neutral 2. negative 3. positive
9. What is the purpose of a salt bridge in an electrochemical cell?
1. To act as a reducing agent 2. To act as an oxidizing agent
3. To provide electrons that flow between the two half cells and complete the electrical circuit
4. To provide ions that flow between the two half cells and complete the electrical circuit
10. What is NOT the same for an electrochemical cell and an electrolytic cell?
A. An oxidation/reduction chemical reaction occurs within them
B. They have two electrodes: a cathode and an anode
C. They produce voltage and an electric current
D. Electrons flow through an external circuit
11. Given the equations A, B, C, and D; which two equations represent redox reactions?
A. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
B. $\text{Cl}_2 + \text{H}_2\text{O} \rightarrow \text{HClO} + \text{HCl}$
C. $\text{CuO} + \text{CO} \rightarrow \text{CO}_2 + \text{Cu}$
D. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

Holt Electrochemistry Review Answers

J Elliott



Holt Electrochemistry Review Answers:

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In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Holt Electrochemistry Review Answers**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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