

Assessment 1**Chapter Test B****Chapter: Chemical Equations and Reactions**

Part 1 In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The precipitation of a slightly soluble solid compound from a solution of two ionic reactants is called the formation of a
- a. gas.
 - precipitate.
 - condensation reaction.
 - halogen.
- _____ 2. To balance a chemical equation, it is permissible to adjust the
- coefficients.
 - subscripts.
 - formula of the products.
 - number of products.
- _____ 3. In a chemical equation, the symbol (aq) indicates that the substance is
- solid.
 - dissolved in water.
 - an acid.
 - molten.
- _____ 4. The tendency for a replacement reaction to occur depends on the
- activity between any two elements in the activity series themselves.
 - temperature difference.
 - valence electrons and mass of.
 - relative between any two elements in the activity series themselves.
- _____ 5. The coefficients in a chemical equation
- indicate the number of moles of each substance that react.
 - show the number of grams of each substance that react.
 - are the molar masses of the substances.
 - show the valence electrons for each atom.
- _____ 6. If metal X is more reactive than metal Y in the activity series, then
- X will replace ions of Y in solution.
 - Y will replace ions of X in solution.
 - Y will react oxidize only indirectly.
 - X will react with cold water.

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KJ Lindholm-Leary

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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Holt Chemistry Ralph Thomas Myers,2004 *Holt Chemistry File* ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J.

Malone,Theodore O. Dolter,2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter *The Development of Modern Chemistry* Aaron J. Ihde,1984-01-01 From ancient Greek theory to the explosive discoveries of the 20th century this authoritative history shows how major chemists their discoveries and political economic and social developments transformed chemistry into a modern science 209 illustrations 14 tables Bibliographies Indices Appendices

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