

## Chapter 6 - Lecture Worksheet 3

Using standard heats of formation, calculate the standard enthalpy change for the following reaction.



$$\begin{aligned}\Delta H^\circ_{\text{rxn}} &= \Delta H^\circ_f(\text{CH}_4, \text{g}) + 2\Delta H^\circ_f(\text{H}_2\text{O}, \text{g}) - \Delta H^\circ_f(\text{CO}_2, \text{g}) - 4\Delta H^\circ_f(\text{H}_2, \text{g}) \\ &= -75 + 2(-242) - (-394) - 4(0) \text{ kJ} = -165 \text{ kJ}\end{aligned}$$

(1) -711 kJ    (2) -165 kJ    (3) -15 kJ    (4) 15 kJ    (5) 165 kJ    (6) 229 kJ    (7) 711 kJ

Using average bond enthalpies estimate the standard enthalpy change for the same reaction.



$$\begin{aligned}\Delta H^\circ_{\text{rxn}} &= 2BE(\text{C}=\text{O}) + 4BE(\text{H}-\text{H}) - 4BE(\text{C}-\text{H}) - 4BE(\text{O}-\text{H}) \\ &\approx 2(728) + 4(436) - 4(413) - 4(463) \text{ kJ} = -304 \text{ kJ}\end{aligned}$$

Which is more accurate ? Why ?

$\Delta H^\circ$  calculated from  $\Delta H^\circ_f$  is more accurate because  $\Delta H^\circ$  calculated from bond enthalpies uses AVERAGE bond enthalpies obtained from a number of different compounds in which the bonds arise and are not specific to the compound in question.

	<u>Standard Enthalpy of formation (kJ/mol)</u>
$\text{CH}_4(\text{g})$	-75
$\text{CO}_2(\text{g})$	-394
$\text{C}_2\text{H}_2(\text{g})$	+227
$\text{H}_2\text{O}(\text{g})$	-242
$\text{H}_2\text{O}(\text{l})$	-286
$\text{NH}_3(\text{g})$	-46
$\text{HCl}(\text{g})$	-92
$\text{HCl}(\text{aq})$	-167
$\text{CaCO}_3(\text{s})$	-1208
$\text{Ca}(\text{OH})_2(\text{aq})$	-1003
$\text{Ca}(\text{OH})_2(\text{s})$	-986
$\text{CaO}(\text{s})$	-635
$\text{SO}_2(\text{g})$	-297
$\text{SO}_3(\text{g})$	-396

	<u>Average Bond Enthalpies (kJ/mol)</u>
H-H	436
N-H	391
N-N	161
N=N	418
N≡N	945
O-H	463
C-O	351
C=O	728
CaO	1072
Cl-Cl	242
Cl-H	432
Br-Br	193
Br-H	366
C-H	413

# Enthalpy Worksheet With Answers

**B Lingard**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## Whispering the Secrets of Language: An Emotional Quest through **Enthalpy Worksheet With Answers**

In a digitally-driven world where screens reign great and quick transmission drowns out the subtleties of language, the profound secrets and emotional nuances concealed within phrases frequently move unheard. However, nestled within the pages of **Enthalpy Worksheet With Answers** a interesting fictional value pulsating with raw feelings, lies a fantastic quest waiting to be undertaken. Penned by a talented wordsmith, that wonderful opus invites viewers on an introspective trip, softly unraveling the veiled truths and profound affect resonating within ab muscles material of each and every word. Within the psychological depths of the poignant evaluation, we shall embark upon a heartfelt exploration of the book is core subjects, dissect their captivating publishing model, and yield to the powerful resonance it evokes strong within the recesses of readers hearts.

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