

23. Limiting Reactant and Excess Reactant

Background

Limiting Reactant and Excess Reactant

In reactions where the reactants are not present in a stoichiometric ratio, the amount of product produced is determined by the amount of the reactant that is fully consumed in the reaction first. This reactant is called the limiting reactant (or limiting reagent). It is the amount of this reactant that is used to calculate the amount of product, just like a baked recipe can't be adjusted based on the amount of an ingredient that's in short supply.

Mass and Molar Ratios in Chemical Reactions

Reactions occur according to the number of atoms and molecules that are available to react in a precise ratio. The law of conservation of mass states that matter can be neither created nor destroyed. Therefore, the number of atoms of each element on the reactants side must be the same as the number of atoms of each element on the products side. This quantitative study of chemical reactions is called stoichiometry. The subscripts in a balanced reaction are related to each other by stoichiometric molar ratios. For example, nitrogen reacts with hydrogen to form ammonia in a 1:3 stoichiometric ratio in the balanced equation:



The molar ratios can be used in calculations that find the amount (in moles) of one substance given the amount of another.

Laboratory Manual Limiting Reactant

Sally A. Henrie



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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures *Laboratory Manual for Principles of General Chemistry* J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section **General, Organic, and Biochemistry Lab Manual** Sara Selfe, 2006-01-12 Offers a choice of classic chemistry experiments and innovative ones All of them place special emphasis on the biological implications of chemical concepts Available for custom publishing at <http://custompub.whfreeman.com> **Green Chemistry Laboratory Manual for General Chemistry** Sally A. Henrie, 2015-03-18 Green chemistry involves designing novel ways to create and synthesize products and implement processes that will eliminate or greatly reduce negative environmental impacts Providing educational laboratory materials that challenge students with the customary topics found in a general chemistry laboratory manual this lab manual enables students to see how green chemistry principles can be applied to real world issues Following a consistent format each lab experiment includes objectives prelab questions and detailed step by step procedures for performing the experiments Additional questions encourage further research about how green chemistry principles compare with traditional more hazardous experimental methods **Laboratory Manual** Jo A. Beran, 1990 Elements of General and Biological Chemistry, Laboratory Manual John R. Holum, 1994-10-19 This updated edition explains recent advances in environmental studies and in the molecular basis

of life Suitable for students interested in the health care field as well as those who want to know how nature and human life work at the molecular level the book begins by providing readers with a solid background in formulas structures equations solutions and equilibria A number of topics are introduced early such as molarity and are discussed in more detail in later chapters Each chapter contains a summary as well as review exercises Practical Lab Manual of Pharmaceutical Organic Chemistry - II Dr. Priyanka Gupta Manglik, 2024-08-15 Builds upon Volume I covering advanced organic chemistry experiments with relevance to drug synthesis mechanisms and qualitative quantitative analysis **Prentice Hall Chemistry**, 2000 *Laboratory Manual of Medicinal Chemistry III* Dr. Niranjan S. Mahajan, Dr. Shrinivas D. Joshi, Dr. T. Prabha, Mr. Shrikrishna S. Maske, We are very pleased to put forth the Laboratory Manual of Medicinal Chemistry III This manual is prepared as per PCI B Pharm course regulations 2014 and is divided into three sections for laboratory techniques determination of oil values and preparations of organic compounds The methods of all the experiments are added from the recent research papers so that the advancement in the methods or apparatus can be addressed This manual is designed for outcome based education and each experiment is arranged in a uniform way such as practical significance practical outcomes ProOs and its mapping with course outcomes theory resources used procedure precautions observations result conclusion references and synopsis questions Each experiment offers an opportunity to perform practical work allowing students to gain proficiency in effectively managing equipment handling glassware chemicals and reagents and writing analytical reports In addition questions are provided at the end of the experiments to enhance students knowledge which will be beneficial for them as they pursue higher studies During the laboratory period you will have to multitask while you are doing experiment It is essential to document properly what you do and what you observe while doing the practical Always plan your work ahead and think about what you are doing why you are doing it what is happening and what you can conclude from your experiment We acknowledge the help and co operation extended by various persons in bringing out this manual We are highly indebted to the authors of various books and articles mentioned in the reference which became a major source of information for writing this manual We also thank the publishers designers and printers who graciously worked hard to publish this manual in time We hope that this manual will assist students in understanding concepts principles and performing procedures We wish you all the best Clean Energy: Hydrogen/fuel Cells Laboratory Manual K S V Santhanam, Gerald A Takacs, Massoud J Miri, Alla V Bailey, Thomas D Allston, Roman J Press, 2016-01-15 This manual is designed for the use of hydrogen as a fuel in the fuel cells The turn of the century has seen a realization of moving towards clean energy due to a variety of considerations ranging from global warming anxiety to living in a healthy atmosphere depletion of fossil fuels oil slick in Gulf of Mexico resulting in disasters and so forth Innumerable debates in the literature has led to the identification of hydrogen as the safest and efficient fuel over the other available fuels This fuel can be used in two ways a direct combustion like gasoline and b fuel cells The use of it by the first method requires pure oxygen to be used for

combustion it is an expensive method involving oxygen storage and transportation. If oxygen is substituted by air in the combustion, it produces nitrogen oxides that are defying the definition of clean energy. The other method is to use it as a fuel cell for easy emission-free transportation. Here, chemical energy is converted to electrical energy directly in a fuel cell. To illustrate principles of related fuel cells, methanol and borohydride fuel cells are included in this manual. The nine experiments described here are designed for illustrating the concepts for the beginners and those motivated to go for clean energy. DVD displays the actual experimental set-up and measurement procedures for Hydrogen safety, Fuel value measurement, Gaseous properties of hydrogen, Proton exchange membrane fuel cell assembly, and Dissolved methanol fuel cell.

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