

23. When red litmus paper is added to lime water  
 (a) it changes to blue  
 (b) it remains red  
 (c) it becomes green  
 (d) it changes into yellow colour.
24. Solution A turns blue litmus red, evolves  $H_2$  gas with Zn metal and gives out  $CO_2$  on treatment with  $Na_2CO_3$ . Solution B turns red litmus blue, liberates  $H_2$  with Zn metal and does not react with  $Na_2CO_3$ . Which of the following option is correct?  
 (a) Solution A is NaOH; solution B is HCl.  
 (b) Solution A is HCl; solution B is NaOH.  
 (c) Solution A is NaOH; solution B is  $Na_2CO_3$ .  
 (d) Solution A is  $Na_2CO_3$ ; solution B is NaOH.
25. A dilute solution of sodium carbonate was added to two test tubes containing dilute HCl (A) and dilute NaOH (B) respectively. The correct observation was  
 (a) a brown coloured gas liberated in test tube A  
 (b) a brown coloured gas liberated in test tube B  
 (c) a colourless gas liberated in test tube A  
 (d) a colourless gas liberated in test tube B.

26. Alkalies are  
 (a) acids dissolved in water  
 (b) bases dissolved in water  
 (c) a mixture of acid and base  
 (d) a diluted solution of HCl.
27. The correct formula for calculation of pH is  
 (a)  $+\log [H^+]$  (b)  $-\log [OH^-]$   
 (c)  $-\log [H^+]$  (d)  $-\log [H^+] + \log [OH^-]$
28. Antacids are taken for \_\_\_\_\_ reaction, when excess acid is formed in stomach.  
 (a) combination (b) decomposition  
 (c) photochemical (d) neutralisation
29. An indicator shows red colour in acid solution and yellow colour in basic solution. The indicator is  
 (a) phenolphthalein (b) methyl orange  
 (c) litmus (d) China rose.
30. The number of ionisable hydrogen ions in  $H_2SO_4$  and HCl are respectively  
 (a) 2, 3 (b) 3, 1  
 (c) 2, 1 (d) 1, 1

### Unboxing



### Apply Your Concept

1. (b) : Sodium carbonate is a basic salt. When dissolved in water, it gives an alkaline (or basic) solution.
2. (d) : The given solution is basic in nature. When excess of HCl is added, it becomes acidic.
3. (a) : The sample Y with  $pH = 13$  is a strongly basic compound. Hence, it is NaOH.
4. (b) : The increasing order of pH is  
 Highly acidic  $\rightarrow$  Moderately acidic  $\rightarrow$  Neutral  $\rightarrow$   
 Moderately basic  $\rightarrow$  Highly basic (alkaline)  
 For the given substances, the increasing order of pH will be HCl,  $CH_3COOH$ , NaOH.
5. (b) : Soft drinks are moderately acidic. In this range of pH, the pH paper will show yellow-orange colour.
6. (b) : In acidic solution, blue litmus paper changes to red and in basic solution, red litmus paper changes to blue. Hence, blue litmus paper first changes its colour to red and then to blue.
7. (a) :  $[H^+] = 1.0$ ,  $pH = -\log [H^+] = -\log (1)$   
 So,  $pH = 0$
8. (c) : The concept of pH is not applicable to dry solid substances. Therefore, pH of solid sodium bicarbonate cannot be found directly using a pH paper.

9. (b) : According to the experiment performed,

| Solutions            | I             | II              | III         | IV                |
|----------------------|---------------|-----------------|-------------|-------------------|
| Colour of pH paper : | Red           | Blue            | Light green | Orange            |
| Conclusion :         | Highly acidic | Highly alkaline | Neutral     | Moderately acidic |

So, I and IV are acidic.

10. (b) : A : Blue Highly alkaline solution Sodium carbonate solution  
 B : Orange Moderately acidic solution Acetic acid
11. (d) : Universal indicators are mixture of many indicators.
12. (d) : The correct method to find pH of a solution is to put the solution on a pH paper using a dropper.
13. (d) : Light green (A) : Almost neutral  
 Dark red (B) : Highly acidic  
 Light orange (C) : Slightly acidic  
 Dark blue (D) : Strong basic  
 Here order of pH:  $D > A > C > B$
14. (a) : Zinc reacts with dil.  $H_2SO_4$  to produce hydrogen gas.  
 $Zn_{(s)} + H_2SO_{(aq)} \rightarrow ZnSO_{(aq)} + H_{2(g)} \uparrow$

# Lab Manual Science Class 10 Cbse Experiment 7

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