

### Pre-lab Questions: The Determination of the Empirical Formula of a Hydrate

A hydrate is an ionic compound (ionic salt) that contains water molecules in its structure. An anhydrate (also called anhydrous salt) is the substance that remains after the water is removed from a hydrate. When a hydrate is heated, the water molecules are removed in the form of steam, leaving behind the anhydrate. To determine the formula of a hydrate experimentally, the mole ratio of the water portion to the mole ratio of the anhydrous portion must be calculated.

#### Questions:

During a Chem 121 experiment, a student obtained the following data after heating a sample of calcium chloride hydrate  $\text{CaCl}_2 \cdot x\text{H}_2\text{O}$ , where  $x$  refers to the number of water molecules hydrated per molecule of  $\text{CaCl}_2$ .

#### Data:

Weight of crucible and lid (g)	32.1070
Weight of crucible, lid and hydrated salt (g)	36.8283
Weight of crucible, lid and anhydrous salt after	
a) 1 <sup>st</sup> heating (g)	35.7347
b) 2 <sup>nd</sup> heating (g)	35.6682
c) 3 <sup>rd</sup> heating (g)	35.6679

Be sure to show the general formula used for each step.

1. Use this information to determine the weight of the hydrated salt ( $\text{CaCl}_2 \cdot x\text{H}_2\text{O}$ ):

a) the weight of the anhydrous salt ( $\text{CaCl}_2$ )

*Note:* the weight after the 2<sup>nd</sup> and 3<sup>rd</sup> heating are close enough in value to assume that constant weight has been reached and all the water molecules have been removed

b) the weight of water removed by heating

# Empirical Formula Determination Pre Lab Question Answers

**DP Hallahan**



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## **Table of Contents Empirical Formula Determination Pre Lab Question Answers**

1. Understanding the eBook Empirical Formula Determination Pre Lab Question Answers
  - The Rise of Digital Reading Empirical Formula Determination Pre Lab Question Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Formula Determination Pre Lab Question Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Empirical Formula Determination Pre Lab Question Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Formula Determination Pre Lab Question Answers
  - Personalized Recommendations
  - Empirical Formula Determination Pre Lab Question Answers User Reviews and Ratings
  - Empirical Formula Determination Pre Lab Question Answers and Bestseller Lists
5. Accessing Empirical Formula Determination Pre Lab Question Answers Free and Paid eBooks
  - Empirical Formula Determination Pre Lab Question Answers Public Domain eBooks
  - Empirical Formula Determination Pre Lab Question Answers eBook Subscription Services
  - Empirical Formula Determination Pre Lab Question Answers Budget-Friendly Options
6. Navigating Empirical Formula Determination Pre Lab Question Answers eBook Formats

- ePub, PDF, MOBI, and More
- Empirical Formula Determination Pre Lab Question Answers Compatibility with Devices
- Empirical Formula Determination Pre Lab Question Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Empirical Formula Determination Pre Lab Question Answers
  - Highlighting and Note-Taking Empirical Formula Determination Pre Lab Question Answers
  - Interactive Elements Empirical Formula Determination Pre Lab Question Answers
- 8. Staying Engaged with Empirical Formula Determination Pre Lab Question Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Empirical Formula Determination Pre Lab Question Answers
- 9. Balancing eBooks and Physical Books Empirical Formula Determination Pre Lab Question Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Empirical Formula Determination Pre Lab Question Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical Formula Determination Pre Lab Question Answers
  - Setting Reading Goals Empirical Formula Determination Pre Lab Question Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical Formula Determination Pre Lab Question Answers
  - Fact-Checking eBook Content of Empirical Formula Determination Pre Lab Question Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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