

EMPIRICAL AND MOLECULAR FORMULAE

EMPIRICAL FORMULA - This represents the simplest whole number ratio of atoms within a molecule

MOLECULAR FORMULA - This represents the exact numbers of each atom in a molecule

Empirical formulae can be calculated from either % mass composition data or mass composition data. Molecular formula can then be calculated if the molecular mass of the compound is known.

EXAMPLE 1 - from % composition

A substance containing carbon and hydrogen and oxygen only contains 73.47% carbon and 10.20% hydrogen by mass, its relative molecular mass is 98. Calculate its molecular formula.

Step 1 calculate empirical formula

ATOMS	CARBON	HYDROGEN	OXYGEN
PERCENTAGE MASS	73.47	10.20	$(100 - (73.47 + 10.20)) = 16.33$
ATOMIC MASS	12	1	16
MASS/ATOMIC MASS	6.1225	10.20	1.020
RATIO(divide by smallest)	6	10	1

Empirical formula = $C_6H_{10}O$

Step 2 use M_r to find molecular formula

$$M_r = \text{mass } (C_6H_{10}O) \times n$$

$$98 = (98) \times n$$

$$n = 1$$

molecular formula =

$C_6H_{10}O$

EXAMPLE 2 - from mass data

A substance contains 12.28g nitrogen, 3.51g hydrogen, 28.07g sulphur and 56.10g oxygen, its M_r is 228. Calculate its molecular formula.

Empirical And Molecular Answers And Questions

L Darling-Hammond



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Table of Contents Empirical And Molecular Answers And Questions

1. Understanding the eBook Empirical And Molecular Answers And Questions
 - The Rise of Digital Reading Empirical And Molecular Answers And Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Answers And Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical And Molecular Answers And Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Answers And Questions
 - Personalized Recommendations
 - Empirical And Molecular Answers And Questions User Reviews and Ratings
 - Empirical And Molecular Answers And Questions and Bestseller Lists
5. Accessing Empirical And Molecular Answers And Questions Free and Paid eBooks

- Empirical And Molecular Answers And Questions Public Domain eBooks
- Empirical And Molecular Answers And Questions eBook Subscription Services
- Empirical And Molecular Answers And Questions Budget-Friendly Options
- 6. Navigating Empirical And Molecular Answers And Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical And Molecular Answers And Questions Compatibility with Devices
 - Empirical And Molecular Answers And Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical And Molecular Answers And Questions
 - Highlighting and Note-Taking Empirical And Molecular Answers And Questions
 - Interactive Elements Empirical And Molecular Answers And Questions
- 8. Staying Engaged with Empirical And Molecular Answers And Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical And Molecular Answers And Questions
- 9. Balancing eBooks and Physical Books Empirical And Molecular Answers And Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical And Molecular Answers And Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical And Molecular Answers And Questions
 - Setting Reading Goals Empirical And Molecular Answers And Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical And Molecular Answers And Questions
 - Fact-Checking eBook Content of Empirical And Molecular Answers And Questions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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