Worksheet: Calculating Empirical & Molecular Formulas

1. The empirical formula for the compound having the formula H₂C₂O₃ is:

ISS CO'H SCI COH [D] CyOuHe

Calculate the empirical formula of a compound that is 85.6% C and 14.4% H (by mass).

BALCH, 181 CH TOT CVH-IDI C.H. (E) C.H.

 A compound is analyzed and found to contain 12.1% carbon, 16.2% oxygen, and 71.7% chlorine (by mass). Calculate the empirical formula of this compound.

C (2.1 - 1 1 U 767 = 2 ([All COCI,] [B] CO₂CI, (C) CO2CI IDI COCH [E] COCK O Thomas

4. A compound contains 40.0% carbon, 6.7% hydrogen, and 53.3% oxygen (by mass). Colculate the empirical formula.

74.04 g 0

 Determine the empirical formula of a compound containing 54.2% F and 45.8% S (by mass).

19.00 g/med = 2.85 med F = 1,99 ~ 2 32-079/mol = 1.43 mol S = 1

T. A compound has 40.68% carbon, 5.12% hydrogen, and 54.20% oxygen (by mass). Calculate its empirical formula.

40.689 C = 3.387 5.129 H = 5.669 54.209 C = 3.387 = 1.496 x2= 3 = 1 ×2 = 2

8. A 7.33-g sample of lanthanum, La, combines with oxygen to give 10.29 g of the oxide. Calculate the empirical formula of this oxide.

10.29g - 7.33g La = 2.96g & 7.339 La = .0027 2.9490 = .1850 = 3.5 x2 = 7

9. Calculate the molecular formula of a compound with the empirical formula CH₂O and a moler mass. of 150 gimel. 150 - 30 = 5 200

[A] C,H,O, [B] C,H,O, [C] C,H,O, EDIT CLINLOW DELICATION OF

Empirical Formula Study Guide With Answer Sheet

Camilla Rothe

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