

1) The empirical formula of a compound represents the lowest whole number mole ratio of component elements.

2) The molecular formula is the actual ratio of elements found with the naturally-occurring sample. This could also equal the empirical formula if the naturally-occurring formula is also in the lowest whole number mole ratio.

3) Molecular = $C_4H_8N_2O_4$ ← each subscript can be reduced by 2

Empirical = $C_2H_4N_2O_2$

4) Compound pair B is the only set of compounds with the same empirical formula as

5) To determine the empirical formula, convert each atom's provided mass into moles and find the lowest whole number mole ratio.

$$0.8007 \text{ g C} \times 1 \text{ mol C} / 12.01 \text{ g C} = 0.066669 \text{ mol C}$$

$$0.9333 \text{ g H} \times 1 \text{ mol H} / 1.01 \text{ g H} = 0.92417 \text{ mol H}$$

$$0.2016 \text{ g H} \times 1 \text{ mol H} / 1.008 \text{ g H} = 0.2 \text{ mol H}$$

moles

$$2.130 \text{ g O} \times 1 \text{ mol O} / 16.00 \text{ g O} = 0.13313 \text{ mol O}$$

divide by the smallest # of

$$0.066669 \text{ mol C} / 0.066617 = 1 \text{ mol C}$$

$$0.92417 \text{ mol H} / 0.066617 = 13.87 \text{ mol H}$$

$$0.2 \text{ mol H} / 0.066617 = 3 \text{ mol H}$$

carbanic acid

$$0.13313 \text{ mol O} / 0.066617 = 2 \text{ mol O}$$

CH_3O_2 = empirical formula for

6) Determine the empirical formula first, by assuming that 100g of the sample is available, allowing for 24.27% C → 24.27g C, etc... Use the relative masses of each element to determine moles and find the lowest whole number mole ratio. Then compare the mass of the empirical formula to that of the mass of the molecular formula provided. This comparison will allow you to determine the molecular formula.

$$24.27\% \text{ C} \rightarrow 24.27 \text{ g C} \times 1 \text{ mol C} / 12.01 \text{ g C} = 2.021 \text{ mol C}$$

$$4.07\% \text{ H} \rightarrow 4.07 \text{ g H} \times 1 \text{ mol H} / 1.008 \text{ g H} = 4.037 \text{ mol H}$$

$$71.65\% \text{ Cl} \rightarrow 71.65 \text{ g Cl} \times 1 \text{ mol Cl} / 35.45 \text{ g Cl} = 2.021 \text{ mol Cl}$$

$$2.021 \text{ mol C} / 2.021 = 1 \text{ mol C}$$

$$4.037 \text{ mol H} / 2.021 = 2 \text{ mol H}$$

$$2.021 \text{ mol Cl} / 2.021 = 1 \text{ mol Cl}$$

CH_2Cl = empirical formula

$$\text{Empirical mass of } CH_2Cl \rightarrow C: 12.01 \text{ g} +$$

$$H: 2(1.008 \text{ g}) +$$

$$Cl: 35.45 \text{ g}$$

$$= 49.476 \text{ g}$$

To compare empirical & molecular mass:

$$\text{Molecular mass} / \text{empirical mass} = n \quad n(\text{empirical formula}) = \text{molecular}$$

formula

$$98.95 \text{ g} / 49.476 \text{ g} = 2$$

$$2(CH_2Cl) = C_2H_4Cl_2$$

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Table of Contents Empirical And Molecular Formulas Answer Key

1. Understanding the eBook Empirical And Molecular Formulas Answer Key
 - The Rise of Digital Reading Empirical And Molecular Formulas Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Formulas Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical And Molecular Formulas Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Formulas Answer Key
 - Personalized Recommendations
 - Empirical And Molecular Formulas Answer Key User Reviews and Ratings
 - Empirical And Molecular Formulas Answer Key and Bestseller Lists
5. Accessing Empirical And Molecular Formulas Answer Key Free and Paid eBooks
 - Empirical And Molecular Formulas Answer Key Public Domain eBooks
 - Empirical And Molecular Formulas Answer Key eBook Subscription Services
 - Empirical And Molecular Formulas Answer Key Budget-Friendly Options
6. Navigating Empirical And Molecular Formulas Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical And Molecular Formulas Answer Key Compatibility with Devices
 - Empirical And Molecular Formulas Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical And Molecular Formulas Answer Key
 - Highlighting and Note-Taking Empirical And Molecular Formulas Answer Key
 - Interactive Elements Empirical And Molecular Formulas Answer Key
8. Staying Engaged with Empirical And Molecular Formulas Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Empirical And Molecular Formulas Answer Key
- 9. Balancing eBooks and Physical Books Empirical And Molecular Formulas Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical And Molecular Formulas Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical And Molecular Formulas Answer Key
 - Setting Reading Goals Empirical And Molecular Formulas Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical And Molecular Formulas Answer Key
 - Fact-Checking eBook Content of Empirical And Molecular Formulas Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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