

Lenovo Tablet 10 User Guide



Lenovo

Manual For Lenovo Tablet

Andrew P Mitchell



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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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oxidation state of nitrogen in the following a no 3 b no 2 c no 2

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni}^{3+} \rightarrow 3 \text{ Ag} + \text{Ni}$ oxidation reduction $25 \text{ Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the

magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent
 $\text{a} \rightarrow \text{i}2\text{o}5$
 $5\text{co} \rightarrow \text{i}2 \rightarrow 5\text{co}2$
 $\text{i}2 \rightarrow \text{o}5$
 $5\text{c} \rightarrow \text{o} \rightarrow \text{i}2 \rightarrow 5\text{c} \rightarrow \text{o}2$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also
 $2\text{sr} \rightarrow \text{o}2$
 $2\text{sro} \rightarrow \text{sr}0$ to $\text{sr}2$ oxidized reducing agent $\text{o}0$ to $\text{o}2$ reduced ox ag
 $2\text{li} \rightarrow \text{s} \rightarrow \text{li}2\text{s}$
 $\text{li}0$ to $\text{li}1$ oxidized red ag
 $\text{s}0$ to $\text{s}2$ reduced ox ag
 $3 \text{2cs} \rightarrow \text{br}2$