

A-LEVEL CHEMISTRY

MARK SCHEME

Practice Paper 1

Maximum marks: 105

View detailed guidance on the conclusions you can draw from your students' performance in these papers on the MERIT welcome page. Understand how your students compare with others and target revision effectively by entering marks into MERIT.

1	(a)	(i)	$[H^+][OH^-]$ OR $[H_3O^+][OH^-]$ <i>Ignore (aq)</i> <i>Must have [] not ()</i>	1
		(ii)	$\sqrt{3.46 \times 10^{-14}}$ ($= 1.86 \times 10^{-7}$) <i>If no square root, CE= 0</i> pH = 6.73 Must be 2dp	2
		(iii)	$[H^+] = 10^{-11.26}$ ($= 4.365 \times 10^{-12}$ OR 4.37×10^{-12}) <i>Mark for working</i> $K_w = [4.365 \times 10^{-12} \text{ OR } 4.37 \times 10^{-12} \times 0.047] = 2.05 \times 10^{-13}$ <i>Allow $2.05 \times 10^{-13} = 2.1 \times 10^{-13}$</i> <i>Mark for answer</i> <i>Ignore units</i>	1
	(b)	(i)	$HCOOH \rightleftharpoons HCOO^- + H^+$ <i>Must have \rightleftharpoons but ignore brackets.</i> OR $HCOOH + H_2O \rightleftharpoons HCOO^- + H_3O^+$ <i>Allow HCO_2^- or $CHOO^-$ ie minus must be on oxygen, so penalise $COOH^-$</i>	1
		(ii)	$K_a = \frac{[H^+][HCOO^-]}{[HCOOH]}$ OR $\frac{[H_3O^+][HCOO^-]}{[HCOOH]}$ <i>Must have all brackets but allow ()</i> <i>Must be $HCOOH$ etc.</i> <i>Allow ecf in formulae from (b)(i)</i>	1
		(iii)	M1 $K_a = \frac{[H^+]^2}{[HCOOH]}$ ($[H^+]^2 = 1.78 \times 10^{-4} \times 0.056 = 9.97 \times 10^{-6}$) <i>Allow HA or HX etc.</i> <i>Allow $[H^+] = \sqrt{(K_a \times [HA])}$ for M1</i>	1
			M2 $[H^+] = 3.16 \times 10^{-3}$ <i>Mark for answer</i>	1
			M3 pH = 2.50 allow more than 2 dp but not fewer <i>Allow correct pH from their wrong $[H^+]$ here only if square root shown but not taken, pH = 5.00 can score max 2 for M1 and M3</i>	1

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Nathan Taber



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