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cesium cs has the following abundance cs 132 20 0 cs 133 75 3 cs 134 4 7 weighted average atomic mass $132 \times 0.200 + 133 \times 0.752 + 134 \times 0.047 = 132.85$ amu

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web in the beanium lab students obtained a sample of beanium a fictitious element in order to learn about isotopes and average atomic mass beanium has three isotopes white brown and speckled which have different masses and different percent abundances

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web 1 obtain a sample of beanium separate the pintonium navium and lentilium and then mass each group record the values in the data table 2 count the number of each isotope by counting the pintonium navium and lentilium 3 calculate the total number of isotopes

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web procedure obtain a sample of the classroom mixture of beanium sort your beanium into its three isotopic bean types count the number of beans in each pile find the mass of each pile of beans determine the average mass of each type of bean based on the samples masses

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web terms in this set 14 what are the objective of the lab calculate from experimental data the atomic mass of the fictitious element beanium by using a weighted average and simple average decide which method gives a better representation of the atomic mass of beanium

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web good luck safety 1 do not eat the beans 2 do not throw the beans 3 do not crush the beans don't mess with the beans pre lab questions 1 there are many beans in your container each bean represents an 2 there are three kinds of beans in your container each kind represents a different

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