

Test for Anions :

Anion	Test	Result
Carbonate (CO_3^{2-})	Add dilute acid (e.g. dilute hydrochloric acid)	Effervescence, CO_2 produced (turns limewater milky)
Sulfite (SO_3^{2-})	Add dilute acid (e.g. dilute hydrochloric acid), warm	SO_2 produced (turns aq. potassium manganate (VII) from purple to colourless)
Chloride (Cl^-) [in solution]	Acidify, then add aq. silver nitrate [Ag^+] Acidify, then add aq. lead nitrate [Pb^{2+}]	White ppt. of AgCl $\text{Ag}^+ (\text{aq}) + \text{Cl}^- (\text{aq}) \rightarrow \text{AgCl} (\text{s})$ White ppt. of PbCl_2
Bromide (Br^-) [in solution]	Acidify, then add aq. silver nitrate	Cream ppt. of AgBr
Iodide (I^-) [in solution]	Acidify, then add aq. silver nitrate (or lead nitrate)	Yellow ppt. of AgI (or PbI_2)
Sulfate (SO_4^{2-}) [in solution]	Acidify, then add aq. barium nitrate (or barium chloride) [Ba^{2+}]	White ppt. of BaSO_4 $\text{Ba}^{2+} + \text{SO}_4^{2-} \rightarrow \text{BaSO}_4$
Nitrate (NO_3^-)	Add aq. sodium hydroxide, then little aluminium (foil) and warm carefully	Ammonia produced (turns damp red litmus paper or universal indicator paper blue)

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1. Understanding the eBook June 2002 Paper 6 Chemistry Igcse
 - The Rise of Digital Reading June 2002 Paper 6 Chemistry Igcse
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2002 Paper 6 Chemistry Igcse
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2002 Paper 6 Chemistry Igcse
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2002 Paper 6 Chemistry Igcse

- Personalized Recommendations
- June 2002 Paper 6 Chemistry Igcse User Reviews and Ratings
- June 2002 Paper 6 Chemistry Igcse and Bestseller Lists
- 5. Accessing June 2002 Paper 6 Chemistry Igcse Free and Paid eBooks
 - June 2002 Paper 6 Chemistry Igcse Public Domain eBooks
 - June 2002 Paper 6 Chemistry Igcse eBook Subscription Services
 - June 2002 Paper 6 Chemistry Igcse Budget-Friendly Options
- 6. Navigating June 2002 Paper 6 Chemistry Igcse eBook Formats
 - ePub, PDF, MOBI, and More
 - June 2002 Paper 6 Chemistry Igcse Compatibility with Devices
 - June 2002 Paper 6 Chemistry Igcse Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2002 Paper 6 Chemistry Igcse
 - Highlighting and Note-Taking June 2002 Paper 6 Chemistry Igcse
 - Interactive Elements June 2002 Paper 6 Chemistry Igcse
- 8. Staying Engaged with June 2002 Paper 6 Chemistry Igcse
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2002 Paper 6 Chemistry Igcse
- 9. Balancing eBooks and Physical Books June 2002 Paper 6 Chemistry Igcse
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2002 Paper 6 Chemistry Igcse
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine June 2002 Paper 6 Chemistry Igcse
 - Setting Reading Goals June 2002 Paper 6 Chemistry Igcse
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2002 Paper 6 Chemistry Igcse

- Fact-Checking eBook Content of June 2002 Paper 6 Chemistry Igcse
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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