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inspector and instructor He is a graduate of Texas State University and holds a Bachelor of Science Degree in Occupational Education A certified instructor of electrical trades he has been awarded a lifetime teaching certificate from the Texas Education Agency in the field of Vocational Education Mr Holder has taught thousands of students at Austin Community College Austin Texas Odessa College at Odessa Texas Technical Vocational Institute of Albuquerque New Mexico Howard College at San Angelo Texas and in the public school systems in Fort Worth and San Antonio Texas He is currently Director of Education for Electrical Seminars Inc of San Marcos Texas Mr Holder is an active member of the National Fire Protection Association International Association of Electrical Inspectors and the International Brotherhood of Electrical Workers

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mole moles of H_2O 40.1822 moles therefore according to the equation

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time the total pressure $p_a \times a$ mole fraction $a = \frac{\text{mole } a}{\text{moles total}}$ $\sum n_i$ all x_i n_i mole fraction of substance
 a is often denoted by x_a and

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by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction
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mass C_6H_6 128.17 g/mol hence C_6H_6 s number of moles $\frac{30}{128.17} = 0.234$ mol molar mass of carbon tet CCl_4 153.84 g/mol

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web since the mole fraction of the O_2 is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature

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