TSBIE - MODEL PAPER CHEMISTRY - I

Time: 3 Hours

Max.Marks: 60

Each Question Carries TWO marks.

 $2 \times 10 = 20$

SECTION - A

- What is Lanthanide Contraction? Give one of its consequences.
- Which of the two ions Ca²⁺ or Zn²⁺ is more stable and why?
- What is Boltzman's constant? Give its value.
- 4 Calculate the weight of 0.1 mole of sodium carbonate.
- 5. What is homogenous equilibrium? Write two homogenous reactions.
- Lithium salts are mostly hydrated. Why?
- 7. What happens when magnesium metal is burnt in air?
- 8 Diamond has high melting point explain.
- What is the effect of water on tin (Sn)?
- Write the structure of a) Neopentane b) Trichloro ethanoic acid.

SECTION - B

Each Question Carries FOUR marks.

 $6 \times 4 = 24$

- Explain the differences between emission and absorption spectrum.
- Balance the following reaction by ion electron method.

 $Cr_2O_7^{-2} + SO_3^{-2} \xrightarrow{H^+} Cr^{3+} + SO_4^{-2}$

- Write the postulates of Kinetic Molecular theory of gases.
- Derive Ideal gas equation.
- 15. Chemical analysis of a carbon compound gave the following percentage composition by weight of the elements present, carbon = 10.06%, hydrogen = 0.84%, chlorine = 89.10%, Calculate the empirical formula of the compound.
- State and explain Hess's law of constant heat summation.

May 2chemistry First Year Paper

Mark Steyvers

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