

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Student Exploration: Covalent Bonds

**Vocabulary:** covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

1. There are eight markers in a full set, but Flora and Frank each only have seven markers. Flora is missing the red marker, and Frank is missing the blue marker.

What can they do so that each has a full set of markers? **They can share the markers.**

2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set?

**They can give the ones they need to each other.**

### Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the *Covalent Bonds* Gizmo, atoms form bonds in this way.

To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** (▶) to see the electrons orbiting the nucleus of each atom.



1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have? \_\_\_\_\_7\_\_\_\_\_
2. Click **Pause** (⏸). Drag an electron from the left atom to the right atom. Click **Play**.  
What happens? **They share an electron.**
3. Click **Pause**, drag an electron from the right atom to the left, and then click **Play**.  
What happens now? **They share another one.**

## Lesson 20 Unit 4 Covalent Bonds Answers

**Shashanka Rajendrachari, Orhan Uzun**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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