

7 Ionic and Metallic Bonding

Big idea BONDING AND INTERACTIONS

7.1 Ions

For students using the Foundation edition, assign problems 1, 3–5, 7–12, 14, 15, 18–20.

Essential Understanding Ions form when atoms gain or lose valence electrons, becoming electrically charged.

Lesson Summary

Valence Electrons Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation.

- ▶ For a representative element, the group number equals the number of valence electrons the atom contains.
- ▶ An electron dot structure shows the symbol of the element and its valence electrons.
- ▶ Atoms tend to gain or lose the number of electrons that will provide the atom with a noble gas electron configuration.

Formation of Cations Cations are positively charged ions formed when an atom loses one or more valence electrons.

- ▶ Atoms and the cations formed from them have different properties.
- ▶ Elements in Group 1 form cations with a charge of $1+$, and those in Group 2 form cations with a charge of $2+$.
- ▶ Many transition metals form more than one cation and do not follow the octet rule.

Formation of Anions Anions are negatively charged ions formed when an atom gains one or more valence electrons.

- ▶ Commonly, the name of an anion ends in *-ide*.
- ▶ Anions form from nonmetallic elements.
- ▶ The anions formed from halogens are known as halides.

After reading Lesson 7.1, answer the following questions.

Valence Electrons

1. What are valence electrons?

Valence electrons are the electrons in the highest occupied energy level of an element's atoms.

2. The valence electrons largely determine the *chemical properties* of an element and are usually the only electrons used in *chemical bonds*.

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L Darling-Hammond



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Table of Contents Lesson 20 1 Reading And Study Workbook

1. Understanding the eBook Lesson 20 1 Reading And Study Workbook
 - The Rise of Digital Reading Lesson 20 1 Reading And Study Workbook
 - Advantages of eBooks Over Traditional Books
2. Identifying Lesson 20 1 Reading And Study Workbook
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lesson 20 1 Reading And Study Workbook
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson 20 1 Reading And Study Workbook
 - Personalized Recommendations
 - Lesson 20 1 Reading And Study Workbook User Reviews and Ratings
 - Lesson 20 1 Reading And Study Workbook and Bestseller Lists

5. Accessing Lesson 20 1 Reading And Study Workbook Free and Paid eBooks
 - Lesson 20 1 Reading And Study Workbook Public Domain eBooks
 - Lesson 20 1 Reading And Study Workbook eBook Subscription Services
 - Lesson 20 1 Reading And Study Workbook Budget-Friendly Options
6. Navigating Lesson 20 1 Reading And Study Workbook eBook Formats
 - ePub, PDF, MOBI, and More
 - Lesson 20 1 Reading And Study Workbook Compatibility with Devices
 - Lesson 20 1 Reading And Study Workbook Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lesson 20 1 Reading And Study Workbook
 - Highlighting and Note-Taking Lesson 20 1 Reading And Study Workbook
 - Interactive Elements Lesson 20 1 Reading And Study Workbook
8. Staying Engaged with Lesson 20 1 Reading And Study Workbook
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lesson 20 1 Reading And Study Workbook
9. Balancing eBooks and Physical Books Lesson 20 1 Reading And Study Workbook
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lesson 20 1 Reading And Study Workbook
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lesson 20 1 Reading And Study Workbook
 - Setting Reading Goals Lesson 20 1 Reading And Study Workbook
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lesson 20 1 Reading And Study Workbook
 - Fact-Checking eBook Content of Lesson 20 1 Reading And Study Workbook
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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