

**KEEP IT SIMPLE SCIENCE***Chemistry Module 5*

# **Chemical Equilibrium**

## **WORKSHEETS**

**Worksheet 1****The Basics**

Student Name.....

1.  
a) When any chemical reaction reaches an equilibrium in a closed system, what becomes true about the concentration of all the chemical species present?

b) Explain why this is so, in the case of a:  
i) static equilibrium.

ii) dynamic equilibrium.

2.  
What are the thermodynamic reasons why some reactions (eg combustion) are never reversible?

3.  
What is meant by "open" or "closed" chemical systems?

4.  
Why can a dynamic equilibrium NOT occur in an open system?

5.  
Write an ionic equation to represent the dynamic equilibrium reached between the solid and the dissolved ions in each of these species, when a solution reaches saturation.

a) potassium bromide

b) magnesium chloride

c) iron (III) nitrate

d) ammonium sulfate

e) calcium carbonate

# Keep It Simple Science Chemistry Answers

**Roger Frost**



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