Anion	Test	Result
Carbonate (CO <sub>3</sub> 2-)	Add dilute acid (e.g. dilute hydrochloric acid)	Effervescence, CO <sub>2</sub> produced (turns limewater milky)
Sulfite (SO₃²·)	Add dilute acid (e.g. dilute hydrochloric acid), warm	SO <sub>2</sub> produced (turns aq. potassium manganate (VII) from purple to colourless
Choride (Cl ) (in solution)	Acidify, then add aq. silver nitrate [Ag*] Acidify, then add aq. lead nitrate [Pb2*]	White ppt. of AgCl Ag* (aq) + Cl* (aq) → AgCl (s) White ppt. of PbCl <sub>2</sub>
Bromide (Br ) [in solution]	Acidify, then add aq. silver nitrate	Cream ppt. of AgBr
lodide (I <sup>-</sup> ) [in solution]	Acidify, then add aq. silver nitrate (or lead nitrate)	Yellow ppt. of Agl (or Pbl <sub>2</sub> )
Sulfate (SO <sub>4</sub> 2-) [in solution]	Acidify, then add aq. barium nitrate (or barium chloride) [Ba <sup>2+</sup> ]	White ppt. of BaSO₄ Ba <sup>2+</sup> + SO₄ <sup>2-</sup> → BaSO₄
Nitrate (NO <sub>3</sub> )	Add aq. sodium hydroxide, then little aluminium (foil) and warm carefully	Ammonia produced (turns damp red litmus paper or universal indicator paper blue)

# June 2002 Paper 6 Chemistry Igcse

Zhe Xu

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