

11 Which compound has the greatest percent composition by mass of sulfur?

- way #64  
(1) BaS (3) MgS  
(2) CaS (4) SrS

way #43  
(BA)  $2\text{HBr} \rightarrow \text{H}_2 + \text{Br}_2$  (RF)  
12 Two molecules of HBr collide and then form  $\text{H}_2$  and  $\text{Br}_2$ . During the collision, the bonds in the HBr molecules are

- (1) broken as energy is absorbed  
(2) broken as energy is released  
(3) formed as energy is absorbed  
(4) formed as energy is released

way #49  
13 Which atom in the ground state has a stable electron configuration?

- (1) carbon (3) neon  
(2) magnesium (4) oxygen

way #46  
14 Which statement describes a multiple covalent bond?

- (1) Two electrons are shared.  
(2) Four electrons are shared.  
(3) Two electrons are transferred.  
(4) Four electrons are transferred.

way #51  
15 The electronegativity difference between the atoms in a molecule of HCl can be used to determine

- (1) the entropy of the atoms  
(2) the atomic number of the atoms  
(3) the first ionization energy of the atoms  
(4) the polarity of the bond between the two atoms

way #5  
16 Which two gases can not be broken down by chemical means?

- (1) CO and He (3) Xe and He  
(2) CO and  $\text{NH}_3$  (4) Xe and  $\text{NH}_3$

17 Two substances in a mixture differ in density and particle size. These properties can be used to

- (1) separate the substances  
(2) chemically combine the substances  
(3) determine the freezing point of the mixture  
(4) predict the electrical conductivity of the mixture

18 Which unit is used to express an amount of thermal energy?

- (1) gram (3) joule  
(2) mole (4) pascal

19 Under which conditions of temperature and pressure does a real gas behave most like an ideal gas?

- (1) low temperature and low pressure  
(2) low temperature and high pressure  
(3) high temperature and low pressure  
(4) high temperature and high pressure

20 According to the kinetic molecular theory for an ideal gas, all gas particles

- (1) are in random, constant, straight-line motion  
(2) are separated by very small distances relative to their sizes  
(3) have strong intermolecular forces  
(4) have collisions that decrease the total energy of the system

21 Which mathematical expression represents the heat of reaction for a chemical reaction?

- (1) (the heat of fusion) - (the heat of vaporization)  
(2) (the heat of vaporization) - (the heat of fusion)  
(3) (the potential energy of the products) - (the potential energy of the reactants)  
(4) (the potential energy of the reactants) - (the potential energy of the products)

22 At 101.3 kPa and 298 K, a 1.0-mole sample of which compound absorbs the greatest amount of heat as the entire sample dissolves in water?

- (1) LiBr -48.83 (3) NaOH -46.51  
(2) NaCl +2.38 (4)  $\text{NH}_4\text{Cl} +14.8$

23 For a reaction at equilibrium, which change can increase the rates of the forward and reverse reactions?

- (1) a decrease in the concentration of the reactants  
(2) a decrease in the surface area of the products  
(3) an increase in the temperature of the system  
(4) an increase in the activation energy of the forward reaction

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## **Table of Contents June 06 Physics Regents Answers Explained**

1. Understanding the eBook June 06 Physics Regents Answers Explained
  - The Rise of Digital Reading June 06 Physics Regents Answers Explained
  - Advantages of eBooks Over Traditional Books
2. Identifying June 06 Physics Regents Answers Explained
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an June 06 Physics Regents Answers Explained
  - User-Friendly Interface
4. Exploring eBook Recommendations from June 06 Physics Regents Answers Explained
  - Personalized Recommendations
  - June 06 Physics Regents Answers Explained User Reviews and Ratings
  - June 06 Physics Regents Answers Explained and Bestseller Lists
5. Accessing June 06 Physics Regents Answers Explained Free and Paid eBooks
  - June 06 Physics Regents Answers Explained Public Domain eBooks
  - June 06 Physics Regents Answers Explained eBook Subscription Services
  - June 06 Physics Regents Answers Explained Budget-Friendly Options
6. Navigating June 06 Physics Regents Answers Explained eBook Formats

- ePub, PDF, MOBI, and More
- June 06 Physics Regents Answers Explained Compatibility with Devices
- June 06 Physics Regents Answers Explained Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of June 06 Physics Regents Answers Explained
  - Highlighting and Note-Taking June 06 Physics Regents Answers Explained
  - Interactive Elements June 06 Physics Regents Answers Explained
- 8. Staying Engaged with June 06 Physics Regents Answers Explained
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers June 06 Physics Regents Answers Explained
- 9. Balancing eBooks and Physical Books June 06 Physics Regents Answers Explained
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection June 06 Physics Regents Answers Explained
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine June 06 Physics Regents Answers Explained
  - Setting Reading Goals June 06 Physics Regents Answers Explained
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 06 Physics Regents Answers Explained
  - Fact-Checking eBook Content of June 06 Physics Regents Answers Explained
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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