

AQA Chemistry

GCSE Exam-style mark scheme

C4 Chemical calculations

100 marks (20%)

1 hour

Question	Answer	Mark	Mark	Mark
10.1.1	1.4 g		1	100
10.1.2	1.48 g	<p>Amount of magnesium carbonate $1.48 \text{ g} \div 142.04 \text{ g mol}^{-1}$ $= 0.01042 \text{ mol}$ $\times 1000 = 10.42 \text{ mmol}$ <u>10</u> mol as amount of Ca^{2+} ions Bromine ions should be provided by 2 moles</p>	10	100
10.1.3	<p>any 1000 factor</p> <ul style="list-style-type: none"> weighing errors lost or incomplete reaction size of magnesium oxide size of magnesium oxide 		10	100
10.2	<p>all calculations must be shown calculate all ions in the same compound as well as the whole mass of the compound as well as the whole mass</p>	<p>Amount of magnesium carbonate $1.48 \text{ g} \div 142.04 \text{ g mol}^{-1}$ $= 0.01042 \text{ mol}$ $\times 1000 = 10.42 \text{ mmol}$ $\times 2 = 20.84 \text{ mmol}$ $\times 1000 = 20.84 \text{ mmol}$ $\times 160.00 \text{ g mol}^{-1}$ $= 3.3344 \text{ g}$ $\approx 3.33 \text{ g}$</p>	10	100
10.3.1	<p>all calculations must be shown</p>		10	100
10.3.2	<p>all calculations must be shown</p>		10	100
10.3.3	<p>all calculations must be shown</p>		10	100

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