

Board Question Paper: March 2014

Chemistry

Time: 3 Hours

Total Marks: 70

Note:

- All questions are compulsory.
- Answer to the two sections are to be written in the same answer book.
- Figure to the right hand side indicate full marks.
- Write balanced chemical equations and draw neat and labelled diagrams wherever necessary.
- Every new question must be started on a new page.
- Use of logarithmic table is allowed

SECTION – I

Q.1. Answer any ONE of the following:

[7]

- What is 'boiling point'?
Derive a relation between ΔH and ΔU for a chemical reaction.
Draw neat labelled diagram of calomel electrode.
Resistance and conductivity of a cell containing 0.001 M KCl solution at 298 K are 1500 Ω and 1.46×10^{-4} S. cm^{-1} respectively. What is the cell constant?
- Write molecularity of the following reaction:
 $2\text{NO}_{(\text{g})} + \text{O}_{2(\text{g})} \longrightarrow 2\text{NO}_{2(\text{g})}$
What is 'calcination'? How does it differ from 'roasting'?
Write resonating structures of ozone.
The decomposition of $\text{N}_2\text{O}_{5(\text{g})}$ at 320 K according to the following equation follows first order reaction:
$$\text{N}_2\text{O}_{5(\text{g})} \rightarrow 2\text{NO}_{2(\text{g})} + \frac{1}{2} \text{O}_{2(\text{g})}$$

The initial concentration of $\text{N}_2\text{O}_{5(\text{g})}$ is 1.24×10^{-2} mol. L^{-1} and after 60 minutes, 0.20×10^{-2} mol. L^{-1} . Calculate the rate constant of the reaction at 320 K.

Q.2. Answer any THREE of the following:

[9]

- One mole of a gas expands by 3 L against a constant pressure of 3 atmosphere. Calculate the work done in:
 - L. atmosphere
 - Joules
 - Calories
- Calculate the amount of CaCl_2 (van't Hoff factor $i = 2.47$) dissolved in 2.5 L solution so that its osmotic pressure at 300 K is 0.75 atmosphere.
Given: Molar mass of CaCl_2 is 111 g. mol^{-1} .
 $R = 0.082$ L. atm. $\text{K}^{-1} \text{mol}^{-1}$
- Describe anomalous behaviour of fluorine with the other elements of group 17 with reference to:
 - Hydrogen bonding
 - Oxidation state
 - Polyhalide ions

Hsc Chemistry 2014 Paper Topic Marking

Clifford Lane



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Table of Contents Hsc Chemistry 2014 Paper Topic Marking

1. Understanding the eBook Hsc Chemistry 2014 Paper Topic Marking
 - The Rise of Digital Reading Hsc Chemistry 2014 Paper Topic Marking
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry 2014 Paper Topic Marking
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry 2014 Paper Topic Marking
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry 2014 Paper Topic Marking
 - Personalized Recommendations
 - Hsc Chemistry 2014 Paper Topic Marking User Reviews and Ratings
 - Hsc Chemistry 2014 Paper Topic Marking and Bestseller Lists
5. Accessing Hsc Chemistry 2014 Paper Topic Marking Free and Paid eBooks

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 - Hsc Chemistry 2014 Paper Topic Marking eBook Subscription Services
 - Hsc Chemistry 2014 Paper Topic Marking Budget-Friendly Options
6. Navigating Hsc Chemistry 2014 Paper Topic Marking eBook Formats
- ePub, PDF, MOBI, and More
 - Hsc Chemistry 2014 Paper Topic Marking Compatibility with Devices
 - Hsc Chemistry 2014 Paper Topic Marking Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Hsc Chemistry 2014 Paper Topic Marking
 - Highlighting and Note-Taking Hsc Chemistry 2014 Paper Topic Marking
 - Interactive Elements Hsc Chemistry 2014 Paper Topic Marking
8. Staying Engaged with Hsc Chemistry 2014 Paper Topic Marking
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Chemistry 2014 Paper Topic Marking
9. Balancing eBooks and Physical Books Hsc Chemistry 2014 Paper Topic Marking
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry 2014 Paper Topic Marking
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hsc Chemistry 2014 Paper Topic Marking
- Setting Reading Goals Hsc Chemistry 2014 Paper Topic Marking
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hsc Chemistry 2014 Paper Topic Marking
- Fact-Checking eBook Content of Hsc Chemistry 2014 Paper Topic Marking
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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