

3.50 To solve this type of problem, read the problem carefully, word for word. Then try to interpret parts of the problem. For example:

- 1) Formaldehyde is an aldehyde, $\text{H}-\overset{\text{O}}{\underset{\text{H}}{\text{C}}}-\text{H}$
- 2) Trioxane - that is, 3 formaldehydes come together to form a compound $\text{C}_3\text{H}_6\text{O}_3$. Because no atoms are eliminated, all of the original atoms are still present.
- 3) There are no carbonyl groups. This means that trioxane cannot contain any $\text{C}=\text{O}$ functional groups. If you look back to Table 3.1, you can see that the only oxygen-containing functional groups that can be present are either ethers or alcohols.
- 4) A monosubstituted derivative is a compound in which one of the H 's has been replaced by a Br . Because only one monosubstituted derivative is possible, we know that there can only be one type of hydrogen in trioxane. The only possibility for trioxane is:



3.51 The highest energy conformation of hexamethane has a strain energy of 15 kJ/mol . Because this includes two $\text{H}-\text{H}$ eclipsing interactions of 4.0 kJ/mol each, the value of an $\text{H}-\text{H}$ eclipsing interaction is $15 \text{ kJ/mol} - 2(4.0 \text{ kJ/mol}) = 7.0 \text{ kJ/mol}$.

Energy

60°

15 kJ/mol



Mcmurry Organic Chemistry 8e Solutions Manual

TD Snyder



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