

Name _____ Class _____ Date _____

Assessment

Chapter Test B

Atomic Physics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. What is the frequency of a photon with an energy of $1.00 \times 10^{-19} \text{ J}$?
 $(h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s})$
 - a. $1.00 \times 10^{14} \text{ Hz}$
 - b. $2.00 \times 10^{14} \text{ Hz}$
 - c. $3.00 \times 10^{14} \text{ Hz}$
 - d. $4.00 \times 10^{14} \text{ Hz}$

- _____ 2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
 - a. 1 : 1
 - b. 2 : 1
 - c. 3 : 1
 - d. 4 : 1

- _____ 3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
 - a. 4.6 eV
 - b. 4.8 eV
 - c. 1.4 eV
 - d. 2.4 eV

- _____ 4. What causes the bright lines in the emission spectrum of an element to occur?
 - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

- _____ 5. What causes the dark lines in the absorption spectrum of an element to occur?
 - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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Clifford Lane



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