



- The nuclear symbol tells you an atom's atomic and mass number
- ATOMIC NUMBER** - amount of protons & electrons
- MASS NUMBER** - total amount of protons & electrons

An **atom** is a really tiny particle. It has a radius of about 0.1 nanometres ( $1 \times 10^{-10}$  m). Atoms are neutral because the no. protons = no. electrons. An atom consists of:

Particle	Relative mass	Charge
Proton	1	+1
Neutron	1	0
Electron	0 (very small)	-1

### THE NUCLEUS

- Middle of the atom
- Contains protons and neutrons
- Has a radius of around  $1 \times 10^{-14}$  m ( $1/100^{\text{th}}$  of atom radius)
- Positively charged due to protons
- Almost whole mass of atom concentrated in nucleus

### THE ELECTRONS

- Move around nucleus in electron shells
- Negatively charged & tiny, but cover a lot of space
- The volume of their orbits determines the size of the atoms
- Negatively charged
- Bond to other electrons to make chemical bond



You can find the number of neutrons by subtracting the mass number from the atomic number  
**Number of neutrons = atomic number - mass number**

**Element** A substance made up of atoms that all have the same number of protons in their nucleus (i.e. all are the same atom)

To balance an equation, you should write out all the elements in the middle, write out how many there are on either side and add numbers to balance it.

**Compound** Substances formed from 2/more elements, where the atoms of each are held together by chemical bonds

**Isotopes** Different forms of the same element, which have the same number of protons but a different number of neutrons.

## CHEMISTRY C1

Atomic structure, separating mixtures

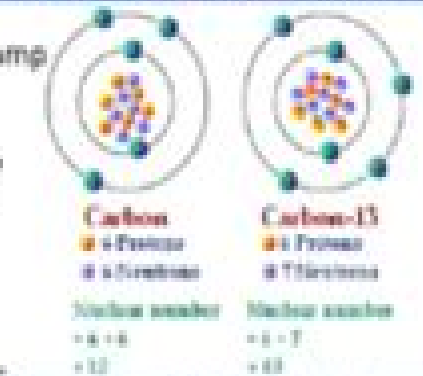


- Making bond involves atoms giving away, taking or sharing electrons
- A chemical bond is required to break these bonds and separate the compound
- There are 2 types of compound bonds:

2 common

- Carbon-12
- Carbon-13

Because many elements can exist as a number of different isotopes, relative atomic mass ( $A_r$ ) is used instead of mass number when referring to the element



$$\text{relative atomic mass} = \frac{\text{sum of (isotope abundance} \times \text{isotope mass number)}}{\text{sum of all isotope abundances}}$$

**EQ** Copper has two stable isotopes. Cu-63 has an abundance of 69.2% and Cu-65 has an abundance of 30.8%. Calculate the relative atomic mass of copper to 1dp.

**A)**  $\frac{(69.2 \times 63) + (30.8 \times 65)}{100} = \frac{4353.6 + 2002}{100} = \frac{6355.6}{100} = 63.6$

The relative atomic mass is an average mass taking into account the different masses & abundances (amounts) of all the isotopes that make up the element

### 1) IONIC BONDING

- When a metal and a non-metal react to form compound
- The metal loses electrons to form a positive ion
- The non-metal gains electrons to form a negative ion
- The opposite charges of the ions means they're strongly attracted

### 2) COVALENT BONDING

- A compound formed from non-metal consists of molecules
- Each atoms shares electrons with another atom

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