

1. The theory that explains the behavior of gases at the molecular level is called the kinetic theory. Boyle's law refers to the relationship between a theoretical gas's volume and its pressure at constant temperature.
2. Gases deviate most from ideal gas behavior under conditions of very low temperature and very high pressure. The molecules of an ideal gas display no attraction or repulsion for one another. Under ordinary conditions, an ideal gas consists entirely of single atoms which explain why gases are so easily compressed. Ideal gas particles travel in straight lines until they collide with each other or with the walls of their container. The approximate kinetic energy of an ideal gas is proportional to the square of the average speed of the gas. The average kinetic energy of the molecules of an ideal gas is directly proportional to the square of the temperature of the gas.
3. A gas's temperature on the scale of a constant volume gas thermometer relates linearly with the walls of the container. In, the pressure exerted by a gas depends on two factors:
 - a. the number of particles that move
 - b. the weight of the particles that move into the container that the gas exerts
4. To measure gas pressure an instrument called a manometer is used.
5. The earth's atmosphere has weight, which creates atmospheric pressure.
6. The instrument used to measure atmospheric pressure is the barometer.

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N Noddings



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