

Pairs of electrons	BP	LP	Shape	Bond angle	E.g.
2	2	0	Linear	$180^\circ$	$\text{CO}_2$
3	3	0	<u>Trigonal planar</u>	$120^\circ$	$\text{BF}_3$
4	4	0	Tetrahedral	$109.5^\circ$	$\text{CH}_4$
4	3	1	Pyramidal	$107^\circ$	$\text{NH}_3$
4	2	2	Non-linear	$104.5^\circ$	$\text{H}_2\text{O}$
6	6	0	Octahedral	$90^\circ$	$\text{SF}_6$

F321 Ocr 2014

**Mike Smith**



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