

GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on s subshell
- readily react to lose 2 electrons
 - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^-$

Trends down the group:

- 1st ionisation energy decreases
 - Distance between outermost electron and nucleus increases
 - More shielding
 - Counteracts effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
 - Much lower than all transition metals

Physical properties:

- Good conductors of heat and electricity
 - Sea of delocalised electrons
- Form white compounds
 - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are toxic or almost toxic
 - Be compounds show some covalent character because small highly charged cation polarises anion shell of anion
- Oxides / hydroxides are basic
 - React with acids (proton acceptors)
 - Example:
 - BaO
 - $\text{BaO} + \text{H}_2\text{O}$ which is exothermic
- Density = 2.83 g/cm^3
 - Average density of transition metals is about 7 g/cm^3
- Compared to group 2 metals:
 - Harder + denser
 - Higher melting point
 - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
 - Reducing agent - donates electrons (reducing other substances)
 - Oxidising agent - accepts electrons (oxidising itself)
- E.g.
 - $\text{ZnO} + \text{CO} \rightarrow \text{Zn} + \text{CO}_2$
 - $\text{ZnO} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$
 - $\text{CO} + 2\text{e}^- \rightarrow \text{CO}^{2-}$

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AW Rasmussen



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Table of Contents F321 May 2013 Ocr Chemistry Unofficial Paper

1. Understanding the eBook F321 May 2013 Ocr Chemistry Unofficial Paper
 - The Rise of Digital Reading F321 May 2013 Ocr Chemistry Unofficial Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying F321 May 2013 Ocr Chemistry Unofficial Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an F321 May 2013 Ocr Chemistry Unofficial Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from F321 May 2013 Ocr Chemistry Unofficial Paper
 - Personalized Recommendations
 - F321 May 2013 Ocr Chemistry Unofficial Paper User Reviews and Ratings
 - F321 May 2013 Ocr Chemistry Unofficial Paper and Bestseller Lists

5. Accessing F321 May 2013 Ocr Chemistry Unofficial Paper Free and Paid eBooks
 - F321 May 2013 Ocr Chemistry Unofficial Paper Public Domain eBooks
 - F321 May 2013 Ocr Chemistry Unofficial Paper eBook Subscription Services
 - F321 May 2013 Ocr Chemistry Unofficial Paper Budget-Friendly Options
6. Navigating F321 May 2013 Ocr Chemistry Unofficial Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - F321 May 2013 Ocr Chemistry Unofficial Paper Compatibility with Devices
 - F321 May 2013 Ocr Chemistry Unofficial Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of F321 May 2013 Ocr Chemistry Unofficial Paper
 - Highlighting and Note-Taking F321 May 2013 Ocr Chemistry Unofficial Paper
 - Interactive Elements F321 May 2013 Ocr Chemistry Unofficial Paper
8. Staying Engaged with F321 May 2013 Ocr Chemistry Unofficial Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers F321 May 2013 Ocr Chemistry Unofficial Paper
9. Balancing eBooks and Physical Books F321 May 2013 Ocr Chemistry Unofficial Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection F321 May 2013 Ocr Chemistry Unofficial Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine F321 May 2013 Ocr Chemistry Unofficial Paper
 - Setting Reading Goals F321 May 2013 Ocr Chemistry Unofficial Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of F321 May 2013 Ocr Chemistry Unofficial Paper
 - Fact-Checking eBook Content of F321 May 2013 Ocr Chemistry Unofficial Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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