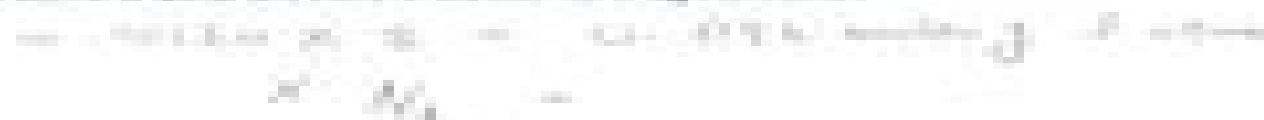


100) Carbonic acid reacts as a buffer at pH 7.4. How many moles of H_2CO_3 react with 1 mole of HCO_3^- to form 1 mole of H_2O and 1 mole of CO_2 ?



Answer: $1 \text{ mole H}_2\text{CO}_3$ reacts with 1 mole HCO_3^- to form $1 \text{ mole H}_2\text{O}$ and 1 mole CO_2 .

101) The only intracellular buffer in body fluid is not the phosphate.

102) Describe how the phosphate buffer works.

The phosphate buffer system is a weak acid (phosphoric acid) and its conjugate base (dihydrogen phosphate). The reaction is as follows:

$$\text{H}_2\text{PO}_4^- \rightleftharpoons \text{HPO}_4^{2-} + \text{H}^+$$

103) Describe how the phosphate buffer works in body fluid.

The phosphate buffer system is a weak acid (phosphoric acid) and its conjugate base (dihydrogen phosphate). The reaction is as follows:

$$\text{H}_2\text{PO}_4^- \rightleftharpoons \text{HPO}_4^{2-} + \text{H}^+$$

104) The phosphate buffer is a component of body fluid buffer system.

105) Describe how the phosphate buffer works in body fluid.

The phosphate buffer system is a weak acid (phosphoric acid) and its conjugate base (dihydrogen phosphate). The reaction is as follows:



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Wolfgang Guggemos



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