

### Types of Chemical Reactions Worksheet #1

#### H/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1)  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  **Synthesis**
- 2)  $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$  **Single Replacement**
- 3)  $3\text{KOH} + \text{AlCl}_3 \rightarrow \text{Al(OH)}_3 + 3\text{KCl}$  **Double Replacement**
- 4)  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  **Synthesis**
- \*5)  $\text{Ca(OH)}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$  **Double Replacement**
- 6)  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$  **Decomposition**
- 7)  $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$  **Single Replacement**
- 8)  $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$  **Double Replacement**
- 9)  $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$  **Synthesis**
- \*10)  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$  **Single Replacement**

\*Remember: "H<sub>2</sub>O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement:  $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement:  $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition:  $\text{Mg(ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement:  $3\text{Ca(OH)}_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe(OH)}_3 + 3\text{CaCl}_2$
- 16) Single replacement:  $\text{Fe} + \text{Cu(NO}_3)_2 \rightarrow \text{Fe(NO}_3)_2 + \text{Cu}$  [Assume Fe<sup>2+</sup>]
- 17) Decomposition:  $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis:  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement:  $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement:  $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$  [Copper (II) is used here]

# Episode 8chemistry Answer Key

**K Morrison**



## **Episode 8chemistry Answer Key:**

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*The Athenaeum* ,1897

**Athenaeum and Literary Chronicle** ,1853

**Athenaeum** ,1863

**Air Pollution**

**Calculations** Daniel A. Vallero, 2023-09-17 Air Pollution Calculations Quantifying Pollutant Formation Transport Transformation Fate and Risks Second Edition enhances the systems science aspects of air pollution including transformation reactions in soil water sediment and biota that contribute to air pollution This second edition will be an update based on research and actions taken since 2019 that affect air pollution calculations including new control technologies emissions measurement and air quality modeling Recent court cases regulatory decisions and advances in technology are discussed and where necessary calculations have been revised to reflect these updates Sections discuss pollutant characterization pollutant transformation and environmental partitioning Air partitioning physical transport of air pollutants air pollution biogeochemistry and thermal reactions are also thoroughly explored The author then carefully examines air pollution risk calculations control technologies and dispersion models The text wraps with discussions of economics and project management reliability and failure and air pollution decision making Provides real life current cases as examples of quantitation of emerging air pollution problems Includes straightforward derivation of equations giving practitioners and instructors a direct link between first principles of science and applications of technologies Presents example calculations that make scientific theory real for the student and practitioner

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