

Here's a list of what has worked for students in the past: 1) review the following list of objectives, 2) answering suggested review problems, 3) read through the book, 4) review old homework problems, 5) read through notes, 6) pay attention during review days, 9) come to class with questions.

When studying, do not make the mistake of telling yourself that you know the material. Convince yourself by working out problems and testing yourself. **Practice, practice, practice!!!**

Also, go to the website and work through the online practice.

**Extra Credit Opportunity:**

(1-10 points) For every **double-sided** sheet of paper that you use to study for the final (working out problems, re-writing notes, correcting test problems), you will receive **one point** credited toward your final exam score (10 point maximum).

**Part I: California State Standards:**

\* standards found on Honor's Exam

**Atomic and Molecular Structure**

1. The periodic table displays the elements in increasing atomic number and shows how periodicity of the physical and chemical properties of the elements relates to atomic structure. As a basis for understanding this concept:
  - a. Students know how to relate the position of an element in the periodic table to its atomic number and atomic mass.
  - b. Students know how to use the periodic table to identify metals, semimetals, nonmetals, and halogens.
  - c. Students know how to use the periodic table to identify alkali metals, alkaline earth metals and transition metals, trends in ionization energy, electronegativity, and the relative sizes of ions and atoms.
  - d. Students know how to use the periodic table to determine the number of electrons available for bonding.
  - e. Students know the nucleus of the atom is much smaller than the atom yet contains most of its mass.
  - g. \* Students know how to relate the position of an element in the periodic table to its quantum electron configuration and to its reactivity with other elements in the table.
  - i. \* Students know the experimental basis for the development of the quantum theory of atomic structure and the historical importance of the Bohr model of the atom.
  - j. \* Students know that spectral lines are the result of transitions of electrons between energy levels and that these lines correspond to photons with a frequency related to the energy spacing between levels by using Planck's relationship ( $E = h\nu$ ).

**Chemical Bonds**

2. Biological, chemical, and physical properties of matter result from the ability of atoms to form bonds from electrostatic forces between electrons and protons and between atoms and molecules. As a basis for understanding this concept:
  - a. Students know atoms combine to form molecules by sharing electrons to form covalent or metallic bonds or by exchanging electrons to form ionic bonds.
  - b. Students know chemical bonds between atoms in molecules such as  $H_2$ ,  $CH_4$ ,  $NH_3$ ,  $H_2CCH_2$ ,  $N_2$ ,  $Cl_2$ , and many large biological molecules are covalent.
  - c. Students know salt crystals, such as  $NaCl$ , are repeating patterns of positive and negative ions held together by electrostatic attraction.

# Mcr Pam Productions Review Packet Semester

**PT Brinkman**



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