

Graphing Exercise #1

Richard C. Weast, ed., "Data for Physics" (Food and Chemistry
Chemical Data, 1970, McGraw-Hill)

Derivation Equation 12.1

$$P/P^0 = \alpha(T)$$

For this experiment for the subject and student, we made a few assumptions about the behavior of the gas. We called these our Kinetic Molecular Theory assumptions:

1. The gas is a perfect gas (ideal gas).
2. The gas is a perfect gas (ideal gas) and is perfectly elastic.
3. The gas is a perfect gas (ideal gas) and is perfectly elastic.

These three assumptions are well-known laws. They don't follow the assumptions that we made for the Kinetic Molecular Theory. The Kinetic Molecular Theory is a gas equation is

$$\left[P - \left(\frac{a}{V^2} \right) \right] (V - b) = nRT$$

The two terms that are subtracted are a/V^2 and b . The term a/V^2 is a function of a and b and is a function of T .

The term b is a function of T and is a function of T . The term a/V^2 is a function of a and b and is a function of T .

The term b is a function of T and is a function of T . The term a/V^2 is a function of a and b and is a function of T .

The term b is a function of T and is a function of T . The term a/V^2 is a function of a and b and is a function of T .

Ideal Gas Graphing Exercise Answers

Christian G. Meyer



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