



Name _____

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Modern Chemistry • CHAPTER 7

HOMEWORK 7-1

(pp. 263–266)

VOCABULARYWrite *true* or *false* for each statement.

1. A chemical formula indicates the relative number of molecules of each kind in a chemical compound.

2. Monoatomic ions are ions formed from a single atom. _____

3. Monoatomic anions of an element are identified by the element's name. _____

4. A binary compound is made of two or more different elements. _____

5. The Stock system is a system for naming chemical elements and their ions. _____

SKILL BUILDER

Explain what each chemical formula tells you about the composition of the compound. The first one is done for you.



The subscript 3 refers to the number of oxygen atoms. There are one atom of sodium and one atom of nitrogen for every three atoms of oxygen.

**STANDARDIZED TEST PREP**

Circle the letter of the best answer.

1. Which statement is true?
- All main-group elements are able to form ions.
 - The total numbers of positive charges and negative charges must be equal in a binary compound.
 - The nonmetals of groups 15, 16, and 17 lose electrons to form cations.
 - Elements in the *d* block form ions of only one charge.
2. Why do many main-group elements lose or gain electrons to form ions?
- The elements are unstable and therefore lose electrons.
 - Their tendency is to maximize potential energy.
 - The elements form covalent bonds, sharing their electrons with other atoms.
 - Their tendency is to form a complete outermost octet in a noble-gas configuration.

Modern Chemistry

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Holt Modern Chemistry Chapter 7 Test Answers

Roman Wölfel



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Table of Contents Holt Modern Chemistry Chapter 7 Test Answers

1. Understanding the eBook Holt Modern Chemistry Chapter 7 Test Answers
 - The Rise of Digital Reading Holt Modern Chemistry Chapter 7 Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Modern Chemistry Chapter 7 Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Modern Chemistry Chapter 7 Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Modern Chemistry Chapter 7 Test Answers
 - Personalized Recommendations
 - Holt Modern Chemistry Chapter 7 Test Answers User Reviews and Ratings

- Holt Modern Chemistry Chapter 7 Test Answers and Bestseller Lists
- 5. Accessing Holt Modern Chemistry Chapter 7 Test Answers Free and Paid eBooks
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 - Holt Modern Chemistry Chapter 7 Test Answers Budget-Friendly Options
- 6. Navigating Holt Modern Chemistry Chapter 7 Test Answers eBook Formats
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 - Holt Modern Chemistry Chapter 7 Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Modern Chemistry Chapter 7 Test Answers
 - Highlighting and Note-Taking Holt Modern Chemistry Chapter 7 Test Answers
 - Interactive Elements Holt Modern Chemistry Chapter 7 Test Answers
- 8. Staying Engaged with Holt Modern Chemistry Chapter 7 Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Modern Chemistry Chapter 7 Test Answers
- 9. Balancing eBooks and Physical Books Holt Modern Chemistry Chapter 7 Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Modern Chemistry Chapter 7 Test Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Holt Modern Chemistry Chapter 7 Test Answers
 - Setting Reading Goals Holt Modern Chemistry Chapter 7 Test Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Holt Modern Chemistry Chapter 7 Test Answers
 - Fact-Checking eBook Content of Holt Modern Chemistry Chapter 7 Test Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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