

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the best (but not all) of the information for this chapter is that.

Focus on this content, but make sure to review other topics: reactions, bonding, molecules, etc.

If you study this document and nothing else, you should at least be able to PASS this test.

Good items will be recall, examples, and/or application of this content. ~~****~~

OBJECTIVES

- Compare with periodic table (modern/transition/lanthanide/actinide) (P 42)
- Compare elements (metals, nonmetals, metalloids) (P 42)
- 5.1: Describe the organization of the periodic table: s-block, p-block, d-block, f-block, and groups (P 42)
- 5.2: Use the Periodic Table to obtain information about the properties (physical and chemical), reactivity, energy, and atomic size of the elements (P 42)
- 5.3: Identify common compounds, metalloids, and noble gases; identify the representative groups and main groups (groups 1, 2, 13, 14, 15, 16, 17, 18) (P 42)
- 5.4: Compare periodic trends of atomic size, ionization energy, and electronegativity (P 42)

5.1: HISTORY OF THE PERIODIC TABLE

• Introductory

- periodic law = the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements
- periodic table = an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column
- lanthanides = a series of the rare earth metals of elements, located below calcium, range from 57 (lanthanum) to 71 (lutetium)
- actinides = set of the metals of heavy radioactive elements that start with 89 (thorium) through 103 (lawrencium) (element number 104) are the periodic table

• Chapter Highlights

- The periodic law states that the physical and chemical properties of the elements, and periodic functions of their atomic numbers,
- The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family
- The columns in the periodic table are referred to as groups

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

• Introductory

- alkali metal = one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
- alkaline-earth metal = one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
- transition elements = set of the metals that start with the inner shell electrons being filled out and so forth
- halogen group elements = the elements in the 17th or 18th of the periodic table
- halogens = set of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts

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M Tight



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