

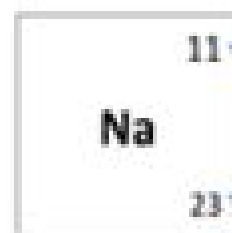
# GCSE

## Additional Science

### **Chemistry (C2)**

#### Revision & Help Cards

#### Back to basics



**Atomic number** = Number of electrons  
= Number of protons

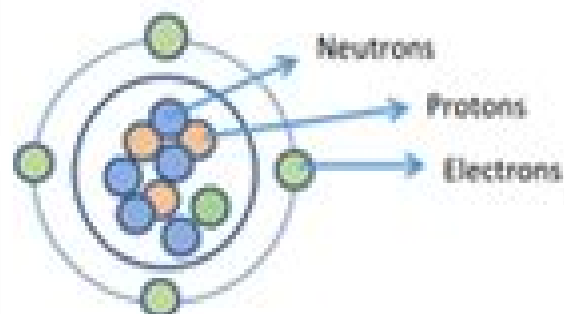
**Mass number** = Number of neutrons +  
Number of protons

To calculate the **number of neutrons** =  
Mass number – Atomic number

**Top Tip** – atomic number is always the **smallest number** but is **NOT** always at the top of the square!

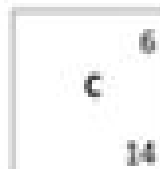
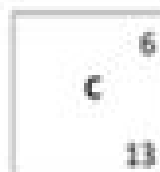
#### Sub-atomic particles

| Particle | Charge | Mass                            |
|----------|--------|---------------------------------|
| Proton   | +1     | 1                               |
| Neutron  | 0      | 1                               |
| Electron | -1     | 1/1836 OR very small/negligible |



**Top Tip** – If there is **more of one type of subatomic particle** in the nucleus, these must be the **neutrons**! Proton number **MUST** equal electron number.

#### Isotopes



| Similarities             | Differences                   |
|--------------------------|-------------------------------|
| Same atomic number       | Different Mass number         |
| Same number of protons   | Different number of neutrons  |
| Same number of electrons | Some isotopes are radioactive |
| Same chemical properties |                               |

**Top Tip** – The relative atomic mass of an element = the **average value** of all the element's isotopes!

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**M Lipman**



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