

## Very LOW concentrations : ppm, ppb & ppt

Certain gases, such as the greenhouse gases (for example, carbon dioxide, water vapour, methane, and ozone), occur in the atmosphere in miniscule amounts. In a random air sample from the troposphere, for example, you would likely find only about 350 molecules of carbon dioxide for every one million molecules of air mixture. Scientists would express this amount as 350 parts per million (ppm).

Gas concentrations can also be expressed in mass units. With gases in the atmosphere, we usually think in terms of volume and may express this as parts per million by volume (ppmv). Some substances occur in such small amounts that scientists measure them in even smaller amounts such as parts per billion by volume (ppbv) or even parts per trillion by volume (pptv).

### **Part per million:**

- 1 can of coke in a 30-metre swimming pool of water
- 1 x 100g chocolate bar on a football pitch

### **Part per billion:**

- 1 bad apple/1,000,000 barrels of apples
- 1 pinch of salt in 10 tons of crisps

### **Part per trillion:**

- 1 flea/360,000,000 elephants

Parts-per million (ppm)	Coefficient / Ratio	Percent (%)	Parts per billion (ppb)	Parts per trillion (ppt)
1 ppm	$1 \times 10^{-6}$	0.0001%	1000 ppb	$1 \times 10^3$ ppt
10 ppm	$1 \times 10^{-5}$	0.001%	10000 ppb	$1 \times 10^4$ ppt
100 ppm	$1 \times 10^{-4}$	0.01%	100000 ppb	$1 \times 10^5$ ppt
1000 ppm	0.001	0.1%	$1 \times 10^5$ ppb	$1 \times 10^6$ ppt
10,000 ppm	0.010	1%	$1 \times 10^7$ ppb	$1 \times 10^8$ ppt
100000 ppm	0.100	10%	$1 \times 10^8$ ppb	$1 \times 10^{11}$ ppt
1,000,000 ppm	1.000	100%	$1 \times 10^9$ ppb	$1 \times 10^{12}$ ppt

# F332 June 2013 Paper

**Chloé Brière**



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