

Selecting TWAIN Settings

The illustration to the right shows the different settings you can adjust in the EPSON TWAIN window, and the preferred order you should adjust them in. Detailed descriptions are given in the electronic *Reference Guide*. Manually adjust settings in the EPSON TWAIN window in the following order:



TIP The most important thing to remember when scanning is to select the correct image type and best resolution for your scan. For more detailed information on making settings in EPSON TWAIN, see the electronic *Reference Guide*.

The screenshot shows the EPSON TWAIN window with the following settings and callouts:

- 1** Select the basic settings: Document Source, Image Type, Destination, Resolution.
- 2** Preview your image.
- 3** Automatically set the scan area for your image using Auto Locate.
- 4** Size or scale the image.
- 5** Customize your settings and configuration, and use color management.
- 6** Scan the image into your application.
- 7** Close the EPSON TWAIN window.

Scanner Light

The scanner has one indicator light. (The indicator light is part of the  Start button.)

The  light indicates the status of the scanner, as described below.

Light	Color	Status
On	Green	The scanner is ready to scan images.
Flashing	Green	The scanner is initializing or is busy scanning.
On	Red	The scanner may not be properly connected to the computer. Unplug the scanner from its power source, remove the USB cable, and then plug them back in. If this does not eliminate the error, one of the following may be the problem and you need to consult your dealer or EPSON®: <ul style="list-style-type: none"> <input type="checkbox"/> The fluorescent lamp needs to be replaced. <input type="checkbox"/> The scanner is malfunctioning.
Off	—	The scanner or AC adapter is unplugged.

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm³ 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less

exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as 57.9 kJ mol⁻¹

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic eg 1 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = -57 \text{ kJ mol}^{-1}$ eg 2 $\text{HNO}_3(\text{aq}) + \text{KOH(aq)} \rightarrow \text{KNO}_3(\text{aq}) + \text{H}_2\text{O(l)}$ for which $\Delta H = -57 \text{ kJ mol}^{-1}$

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(aq)}$ you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts