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R Barnett



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$$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$$

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$$2\text{H}_2\text{S} + 3\text{O}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{SO}_2$$

8 3 mole $\text{H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} \times 18 \text{ g H}_2\text{O} = 149 4 \text{ g H}_2\text{O}$ is your theoretical yield so percent yield
$$\frac{137 1 \text{ g H}_2\text{O}}{149 4 \text{ g H}_2\text{O}} \times 100 = 91 77\%$$

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21 3 grams of O_2 produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have
$$\frac{21 3 \text{ g O}_2}{32 \text{ g O}_2/\text{mole}} = 0 66 \text{ moles O}_2$$

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$$\frac{\text{calculated mass of the product}}{\text{mass of the product expected}} \times 100$$

in this reaction a yield 80 is anticipated

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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