

“What Do you Think?” A PhET Simulation about Molecular Geometry and Polarity.

A study of electronegativity, bond polarity, and molecular polarity (part 1)

Introduction:

In this atomic-level simulation, you will investigate how atoms' electronegativity value affects the bond's behavior. When two atoms bond, a pair of electrons is shared between atoms. Electronegativity is a measure of a single atom's ability to hold electrons, shared in that bond. In this lab you will work diligently, at your own pace, to answer a number of questions.

If a ligand, from what you've already learned about the protons and electrons in an atom, what related issues an atom is likely to have a high electronegativity value? A high number of valence electrons, mean an atom is also small atoms with low shielding.

Why might an atom have a low electronegativity value?

atoms with more than one shell (3, 4, 5) large atoms with lots of shielding.

Procedures:

- Turn on (check) all view options: ORBIT and spin that molecule with your mouse.
- It also turn time and investigate how the binary compound's bond behaves when the atom's electronegativity and oxidation are changed. Do not rush through this step.

Describe the polarity of the bond formed between two atoms with similar, low electronegativity.

_____nonpolar, no molecular dipole_____

Describe the bond formed between two atoms with similar, high electronegativity.

_____nonpolar, no molecular dipole_____

Describe the bond formed between two atoms with very different electronegativity.

_____polar, has molecular dipole_____

Describe (in your own words) what is meant by partial charges, δ^- and δ^+ . _____positive has or negative has_____

δ^- represent to _____partial negative_____

δ^+ represent to _____partial positive_____

What happens when the electric field is applied to a very polar molecule? _____molecule align itself with the field_____

Why do you think this is? _____opposites attract and likes repel_____

What is electron density? _____electron is low is more attracted to one atom or side of a molecule_____

How does the density around a partial positive compare to the density of a partial negative? _____

_____partial positive density < partial negative density_____

What would bring about a higher electron density around an atom? _____higher electronegativity value_____

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SOLUT HONS:**Introduction (10 points)**

Name _____

☐ All of them, even

☐ None of them

☐ All of them

Surface

☐ None

☐ A small amount

☐ A large amount

Electron density

☐ None

☐ A small amount

☐ A large amount

Molecule Polarity Phet Lab Answer Key

PT Brinkman



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