

Key

Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

Model 1 – Lewis Structures

Lewis Structures			
1. H_2CO		H_2CO 3 electron domains (3 bonding, 0 nonbonding)	 Trigonal Planar
2. BeF_2		BeF_2 2 electron domains (2 bonding, 0 nonbonding)	 Linear
3. CH_4		CH_4 4 electron domains (4 bonding, 0 nonbonding)	 Tetrahedral
4. NH_3		NH_3 4 electron domains (3 bonding, 1 nonbonding)	 Pyramidal
5. H_2O		H_2O 4 electron domains (2 bonding, 2 nonbonding)	 Bent
6. CO_2		CO_2 2 electron domains (2 bonding, 0 nonbonding)	 Linear

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