

Name: KEY

Date: \_\_\_\_\_

Complete the table by determining the number of protons, neutrons, electrons, valence electrons & ionic charge.

neutral  
ion  
isotope

	Protons	Neutrons	Electrons	Valence e-	Ionic charge
Cl	17	18	17	7	0
Al <sup>3+</sup>	13	14	10	3	+3
Gallium- 71	31	39	31	3	0

What element has the electron configuration  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^2$ ? Sn - Tin

Write the orbital notation for each of the following elements:

- Magnesium  $1s^2 2s^2 2p^6 3s^2$
- Germanium  $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^{10} 4p^2$

Write the electron configurations for the following elements:

- Mn  $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^5$
- Nb  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^3$

Write the Nobel gas notation for the following elements:

- Cadmium  $[Kr] 5s^2 4d^{10}$
- Tellurium  $[Kr] 5s^2 4d^{10} 5p^4$

Give the quantum numbers for the LAST electron in each element:

- Strontium

$$n = 5$$

$$l = 0$$

$$m_l = 0$$

$$s = +1/2$$

- Nickel

$$n = 4$$

$$l = 2$$

$$m_l = 0$$

$$s = +1/2$$

# Mole Webquest 2 Answer Key

**L Towne**



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